Percent Composition and Molecular Formula Worksheet Key

- What's the empirical formula of a molecule containing 65.5% carbon, 5.5% hydrogen, and 29.0% oxygen? C₃H₃O mass = 55 g/mole
- If the molar mass of the compound in problem 1 is 110 grams/mole, what's the molecular formula?
 C₆H₆O₂
- What's the empirical formula of a molecule containing 18.7% lithium, 16.3% carbon, and 65.0% oxygen? Li₂CO₃
- If the molar mass of the compound in problem 3 is 73.8 grams/mole, what's the molecular formula?
 Li₂CO₃

Write the molecular formulas of the following compounds:

- A compound with an empirical formula of C₂OH₄ and a molar mass of 88 grams per mole. C₄O₂H₈
- 6. A compound with an empirical formula of C4H4O and a molar mass of 136 grams per mole. C8H8O2
- A compound with an empirical formula of CFBrO and a molar mass of 254.7 grams per mole. C₂F₂Br₂O₂
- 8. A compound with an empirical formula of C2H8N and a molar mass of 46 grams per mole. C2H8N

Answer the following questions:

- The percentage composition of acetic acid is found to be 39.9% C, 6.7% H, and 53.4% O. Determine the empirical formula of acetic acid. CH₂O
- 10. The molar mass for question #9 was determined by experiment to be 60.0 g/mol. What is the molecular formula? C₂H₄O₂
- 11. Aniline, a starting material for urethane plastic foams, consists of C, H, and N. Combustion of such compounds yields CO₂, H₂O, and N₂ as products. If the combustion of 9.71 g of aniline yields 6.63 g H₂O and 1.46 g N₂, what is its empirical formula? C₆H₇N
- 12. The molar mass of aniline is 93 g/mol. What is its molecular formula? CoH1N
- Calculate the mass percent of carbon, nitrogen and oxygen in acetamide, C₂H₅NO. %C 40.668 %H 8.533 %N 23.713 %O 27.086
- 14. A 50.51 g sample of a compound made from phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39 g of phosphorus atoms were produced. What is the empirical formula of the compound? PCIs
- When 2.5000 g of an oxide of mercury, (Hg_xO_y) is decomposed into the elements by heating, 2.405 g of mercury are produced. Calculate the empirical formula. Hg₂O
- 16. The compound benzamide has the following percent composition. What is the empirical formula?
 C = 69.40 % H= 5.825 % O = 13.21 % N= 11.57 % C₇H₇NO
- 17. A component of protein called serine has an approximate molar mass of 100 g/mole. If the percent composition is as follows, what is the empirical and molecular formula of serine?

C₃H₂NO₃ empirical formula C₄H₂NO₃ molecular formula

Percent Composition And Molecular Formula Answers

Charles Henrickson

Percent Composition And Molecular Formula Answers:

Counting Moles Nigel P. Freestone, 2016-01-04 Students studying chemistry often struggle with the mole Counting Moles provides an effective aid to learning by giving clear and confident presentation of the essentials of the mole concept needed by those starting chemistry courses This user friendly self teach e book is split into six chapters which sequentially introduce the mole calculating frame to help solve problems Over 200 fully worked examples are given along with several hundred questions The mole concept is applied to topics such as relative atomic mass and relative formula mass percentage composition empirical and molecular formula The book also covers concentration its units volumetric analysis and the relationship between volume mass and moles of gases Counting Moles culminates in you taking a Mole Driving Test On passing this test you are issued with a Counting Moles Driving License that will give you all the confidence required to correctly answer all mole calculations CliffsStudySolver: Chemistry Charles Henrickson, 2007-05-03 The CliffsStudySolver workbooks combine 20 percent review material with 80 percent practice problems and the answers to help make your lessons stick CliffsStudySolver Chemistry is for students who want to reinforce their knowledge with a learn by doing approach Inside you ll get the practice you need to learn Chemistry with problem solving tools such as Clear concise reviews of every topic Practice problems in every chapter with explanations and solutions A diagnostic pretest to assess your current skills A full length exam that adapts to your skill level A glossary examples of calculations and equations and situational tasks can help you practice and understand chemistry This workbook also covers measurement chemical reactions and equations and matter elements compounds and mixtures Explore other aspects of the language including Formulas and ionic compounds Gases and the gas laws Atoms The mole elements and compounds Solutions and solution concentrations Chemical bonding Acids bases and buffers Practice makes perfect and whether you re taking lessons or teaching yourself CliffsStudySolver guides can help you make the grade State Board Ouestions and Answers ,1911 **Fundamentals of Environmental Chemistry, Third Edition** Stanley E. Manahan, 2011-03-05 Written by an expert using the same approach that made the previous two editions so successful Fundamentals of Environmental Chemistry Third Edition expands the scope of book to include the strongly emerging areas broadly described as sustainability science and technology including green chemistry and industrial ecology The new edition includes Increased emphasis on the applied aspects of environmental chemistry Hot topics such as global warming and biomass energy Integration of green chemistry and sustainability concepts throughout the text More and updated questions and answers including some that require Internet research Lecturers Pack on CD ROM with solutions manual PowerPoint presentations and chapter figures available upon qualifying course adoptions The book provides a basic course in chemical science including the fundamentals of organic chemistry and biochemistry The author uses real life examples from environmetral chemistry green chemistry and related areas while maintaining brevity and simplicity in his explanation of concepts Building on this foundation the book covers environmental chemistry broadly

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Applied Chemistry for Polytechnic and Engineering Courses Dr. R.S. Chauhan, 2021-05-27 The book includes the following chapters in details Language of Chemistry Atomic Structure The Periodic table and Atomic properties Water Chemical Bonding Solutions Electrolysis Environmental Chemistry Experiments **CliffsNotes Chemistry Practice Pack** Charles Henrickson, 2010-02-08 About the Contents Pretest Helps you pinpoint where you need the most help Topic Area Reviews Measurement and Units of Measurement Matter Elements Compounds and Mixtures Atoms I The Basics Formulas and Names of Ionic Compounds Acids and Bases The Mole Elements and Compounds Percent Composition and Empirical and Molecular Formulas Chemical Reactions and Chemical Equations Calculations Using Balanced Equations Atoms II Atomic Structure and Periodic Properties Chemical Bonding The Formation of Compounds Gases and the Gas Laws The Forces between Molecules Solids and Liquids Solutions and Solution Composition Acids Bases and Neutralization Glossary Customized Full Length Exam Covers all subject areas Pretest that pinpoints what you need to study most Clear concise reviews of every topic Targeted example problems in every chapter with solutions and explanations Customized full length exam that adapts to your skill level Chemistry Richard Post, Chad Snyder, Clifford C. Houk, 2020-08-17 A practical complete and easy to use guide for understanding major chemistry concepts and terms Master the fundamentals of chemistry with this fast and easy guide Chemistry is a fundamental science that touches all other sciences including biology physics electronics environmental studies astronomy and more Thousands of students have successfully used the previous editions of Chemistry Concepts and Problems A Self Teaching Guide to learn chemistry either independently as a refresher or in parallel with a college chemistry course This newly revised edition includes updates and additions to improve your success in learning chemistry This book uses an interactive self teaching method including frequent questions and study problems increasing both the speed of learning and retention Monitor your progress with self tests and master chemistry quickly This revised Third Edition provides a fresh step by step approach to learning that requires no prerequisites lets you work at your own pace and reinforces what you learn ensuring lifelong mastery Master the science of basic chemistry with this innovative self paced study guide Teach yourself chemistry refresh your knowledge in preparation for medical studies or other coursework or enhance your college chemistry course Use self study features including review questions and quizzes to ensure that you re really learning the material Prepare for a career in the sciences medicine or engineering with the core

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students with a complete course in chemical literacy and assumes minimal prior experience of science and maths Written in an accessible and succinct style this book offers comprehensive coverage of all the core topics in organic inorganic and physical chemistry Topics covered include bonding moles solutions and solubility energy changes equilibrium organic compounds and spectroscopy Each unit contains in text exercises and revision questions to consolidate learning at every step and is richly illustrated with diagrams and images to aid understanding This popular text is an essential resource for students who are looking for an accessible introductory textbook It is also ideal for non specialists on courses such as general science engineering environmental health or life sciences New to this Edition A foreword by Professor Sir John Meurig Thomas FRS former Director of the Royal Institution Three additional units on Gibbs Energy Changes Organic Mechanisms and Fire and Flame Accompanying online resources for this title can be found at bloomsburyonlineresources com chemistry 5e These resources are designed to support teaching and learning when using this textbook and are available at no extra cost

Basic Concepts of Chemistry Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives. These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance Understanding Essential Chemistry Max Diem, 2025-06-11 Enables students to understand apply and retain key concepts in general chemistry Understanding Essential Chemistry offers a unique and approachable supplement to standard general chemistry textbooks designed specifically to aid students in mastering fundamental principles Drawing on extensive classroom experience chemistry professor Max Diem presents key concepts in an uninterrupted flow allowing students to follow a clear and straightforward path to comprehension With a logical algebraic framework the book is structured to build students confidence by breaking down complex topics into manageable pieces and encouraging critical thinking at every step Aimed at STEM majors this book includes checkpoints with example problems and final answers to reinforce concepts and promote independent problem solving skills By methodically emphasizing basic understanding this hands on guide gives students the tools to grasp the core chemistry principles necessary for success in their courses labs and future studies A must have survival guide to boost student confidence in the subject the text Presents chemistry concepts in a streamlined continuous format for easier comprehension and retention Encourages independent critical thinking with targeted example problems with provided solutions Supports any primary general chemistry textbook making it adaptable for various curricula

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