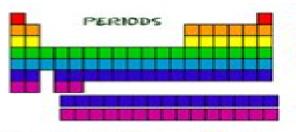
#### Unit 3 Notes: Periodic Table Notes

- John Newlands proposed an organization system based on increasing atomic mass in 1864.
- He noticed that both the chemical and physical properties repeated every 8
  elements and called this the Law of Octaves
- In 1869 both Lothar Meyer and Dmitri Mendeleev showed a connection between atomic mass and an element's properties.
- Mendeleev published first, and is given credit for this.
- He also noticed a periodic pattern when elements were ordered by increasing \_\_\_\_Atomic Mass
- By arranging elements in order of increasing atomic mass into columns,
   Mendeleev created the first Periodic Table.
- This table also predicted the existence and properties of undiscovered elements.
- After many new elements were discovered, it appeared that a number of elements were out of order based on their Properties
- By rearranging the elements based on \_\_\_\_\_\_ Atomic Number \_\_\_\_, the problems with the Periodic Table were corrected.
- This new arrangement creates a periodic repetition of both physical and chemical properties known as the \_\_\_\_\_Periodic Law\_\_\_\_.



Periods are the \_\_\_\_Rows\_\_\_\_

GROUPS

Groups/Families are the Columns

Valence electrons across a period are in the same energy level There are equal numbers of valence electrons in a group.

When elements are arranged in order of increasing <u>Atomic Number</u>,
 there is a periodic repetition of their <u>physical</u> and <u>chemical properties</u>

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## **Table of Contents Periodic Trends Packet Key**

- 1. Understanding the eBook Periodic Trends Packet Key
  - The Rise of Digital Reading Periodic Trends Packet Key
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Periodic Trends Packet Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Trends Packet Key
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Periodic Trends Packet Key
  - Personalized Recommendations
  - Periodic Trends Packet Key User Reviews and Ratings
  - Periodic Trends Packet Key and Bestseller Lists
- 5. Accessing Periodic Trends Packet Key Free and Paid eBooks
  - Periodic Trends Packet Key Public Domain eBooks
  - Periodic Trends Packet Key eBook Subscription Services
  - Periodic Trends Packet Key Budget-Friendly Options

- 6. Navigating Periodic Trends Packet Key eBook Formats
  - o ePub, PDF, MOBI, and More
  - Periodic Trends Packet Key Compatibility with Devices
  - Periodic Trends Packet Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Trends Packet Key
  - Highlighting and Note-Taking Periodic Trends Packet Key
  - Interactive Elements Periodic Trends Packet Key
- 8. Staying Engaged with Periodic Trends Packet Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Trends Packet Key
- 9. Balancing eBooks and Physical Books Periodic Trends Packet Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Trends Packet Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Packet Key
  - Setting Reading Goals Periodic Trends Packet Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Packet Key
  - Fact-Checking eBook Content of Periodic Trends Packet Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements

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