

Periodic and Groups: Valence Electrons: Mastery

Periodic and Groups

Elements that have the same number of valence electrons have similar properties, so we put them in groups. The the periodic table groups + columns that look like this. They will see family groups 1A-18A. Hydrogen is in the same period as Helium.

Elements that have electrons in the outer shell have similar properties in the Periodic Table. Periods go left to right. Elements in the same vertical columns

Period	Group		Group					Group	
	1A	2A	3A	4A	5A	6A	7A	8A	18A
1	H 1H	He 2He	Transition metals (not shown here)					Li 3Li	Ne 10Ne
2	B 5B	C 6C						Be 4Be	Ar 18Ar
3	Al 13Al	Si 14Si						B 5B	Kr 36Kr
4	K 19K	Ca 20Ca						Ge 32Ge	Xe 54Xe

Thinking like the game "Monopoly", answer that an element is in period 3 and group 1.

What element is in Group 1A and Period 2?

What group and period is Helium in?

What element is in Group 2, and Period 18A? (Group 18 is also called the Noble Gas group.)

What element is in Group 2A and Period 2?

What group Helium is in?

Valence Electrons

Valence Electrons are the most important electrons in an atom. Each group contains elements that have the same number of valence electrons. Valence electrons are the electrons that are involved in chemical bonding.

Hydrogen 1



1 valence electron

Helium 2



2 valence electrons

Group 17



7 valence electrons

Group 18, including Helium and the noble gases, have very stable electron shells. How many?

How many electrons?

Hydrogen and Helium are both in Group 1A (hydrogen) 2 (helium).

Group 18: Helium has more stable than any other shell of electrons. Elements in Group 18 and Helium are called noble gases. Helium, neon, argon, krypton, xenon, and radon are noble gases. "Noble gases" usually Group 18A have a full outer shell of valence electrons (usually 8, but Helium has only 2). They are called noble gases because they are very stable.

Periods And Groups Valence Electrons Masses Answers

SJ Ball



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