

Periodic Table Trend Activities:



Purpose: The Periodic Table is a useful tool that helps us predict properties of various elements. In this activity, we will look at 3 properties:

1. **Atomic radius:** The distance from the atomic nucleus to the outmost occupied orbital in a atom
2. **Electronegativity:** The ability of an atom to attract a bonded pair of electrons from another atom
3. **Ionization Energy:** The energy needed to remove an electron from a neutral gaseous atom
4. **Electron Affinity:** The energy released when an electron is added to a neutral gaseous atom

Procedure: Click on the [Cool Periodic Table link](http://www.lynchburg.net/hhs/chemistry/trends/) (click on the element symbol). Record the atomic radius, electronegativity, and first ionization energy for each element in the table below:

<http://www.lynchburg.net/hhs/chemistry/trends/>

Data:

Element	Atomic Radius	Electronegativity	1 st Ionization Energy
Li			
Be			
B			
Na			
Mg			
Al			
C			
Si			
N			
P			

Discussion:

1. What appears to be the trend in atomic radius as you move from left to right in a row?
2. What appears to be the trend in atomic radius as you move down a column?

Periodic Lab Table Trend Activities Answer Key

Katrin Zwirgmaier

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Hands-On General Science Activities With Real-Life Applications Pam Walker, Elaine Wood, 2008-04-21 In this second edition of Hands On General Science Activities with Real Life Applications Pam Walker and Elaine Wood have completely revised and updated their must have resource for science teachers of grades 5 12 The book offers a dynamic collection of classroom ready lessons projects and lab activities that encourage students to integrate basic science concepts and skills into everyday life

The Periodic Table I D. Michael P. Mingos, 2020-02-05 As 2019 has been declared the International Year of the Periodic Table it is appropriate that Structure and Bonding marks this anniversary with two special volumes In 1869 Dmitri Ivanovitch Mendeleev first proposed his periodic table of the elements He is given the major credit for proposing the conceptual framework used by chemists to systematically inter relate the chemical properties of the elements However the concept of periodicity evolved in distinct stages and was the culmination of work by other chemists over several decades For example Newland s Law of Octaves marked an important step in the evolution of the periodic system since it represented the first clear statement that the properties of the elements repeated after intervals of 8 Mendeleev s predictions demonstrated in an impressive manner how the periodic table could be used to predict the occurrence and properties of new elements Not all of his many predictions proved to be valid but the discovery of scandium gallium and germanium represented sufficient vindication of its utility and they cemented its enduring influence Mendeleev s periodic table was based on the atomic weights of the elements and it was another 50 years before Moseley established that it was the atomic number of the elements that was the fundamental parameter and this led to the prediction of further elements Some have suggested that the periodic table is one of the most fruitful ideas in modern science and that it is comparable to Darwin s theory of evolution by natural selection proposed at approximately the same time There is no doubt that the periodic table occupies a central position in chemistry In its modern form it is reproduced in most undergraduate inorganic textbooks and is present in almost every chemistry lecture room and classroom This first volume provides chemists with an account of the historical development of the Periodic Table and an overview of how the Periodic Table has evolved over the last 150 years It also illustrates how it has guided the research programmes of some distinguished chemists

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