

Name: _____

Part 1: Use the periodic table below to answer questions 1-12. Multiple answers may apply.

1. Which letter(s) represent metals? _____
2. Which letter(s) represent nonmetals? _____
3. Which of the following combinations represent an ionic compound? _____
 - a. A & B
 - b. B & E
 - c. A & E
 - d. C & D
 - e. C & F
 - f. D & E
4. Which of the following combinations represent a covalent molecule? _____
 - a. A & C
 - b. D & E
 - c. B & D
 - d. C & F
 - e. E & F
 - f. C & D
5. Which letter(s) represent elements that use roman numerals in their chemical name(s)? _____
6. Which letter(s) will form a cation? _____
7. Which letter(s) will form an anion? _____
8. Which letter(s) will gain three electrons to be stable? _____
9. Which letter(s) will acquire a +3 charge? _____
10. Which letter has the highest ionization energy? _____
11. Which letter has the highest electronegativity? _____
12. Which letter(s) will acquire a positive charge in following the octet rule? _____

Part 2: Answer the questions below

13. How many electrons does Mg^{2+} have? _____
14. How many protons does N^{3-} have? _____
15. What will be the ion symbol (formula), if the ion has 19 protons and 18 electrons? _____
16. How many electrons does an ion have if it has 16 protons and a -2 charge? _____
17. If an ion has 53 protons and 54 electrons, what will be its charge? _____
18. Given the formula $\text{X}(\text{NO}_3)_3$, what is the charge on ion X? Be sure to include the sign (+ or -) . _____
19. How many electrons does the Copper ion have in the ionic compound Cu_2SO_4 ? _____
20. Why do ionic compounds **only** in the liquid state conduct electricity?

Periodic Table And Bonding Study Guide Answers

**T. W. Graham Solomons, Craig B.
Fryhle, Scott A. Snyder**



Periodic Table And Bonding Study Guide Answers:

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