

Trends on the Periodic Table

Part 1: Identifying chemical properties/characteristics of the elements

For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.)

Lithium
Being one of the smaller elements with only 3 electrons, lithium is happy to give one of its electrons away and never seems to ask for any in return.

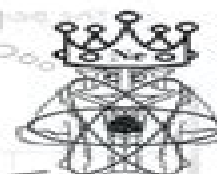


Francium
Francium is the largest of all of the elements and has so many electrons that it has trouble keeping track of them. Francium will eagerly give up electrons to other elements and never takes in extras. But watch out, it has an extremely reactive and explosive personality!

Fluorine
One of the smallest elements, Fluorine is often described as snappy with a "magnetic" personality. As part of the halogen family (stoutest electron thieves), Fluorine is the best at convincing other elements to give up their electrons to him. However, he holds on tight to his own electrons and will not let them go without a fight.



Iodine
As part of the halogen family, Iodine is also a skilled electron thief, however, not nearly as good as its smaller brothers and sisters. Iodine will also put up a pretty good fight to keep its own electrons.

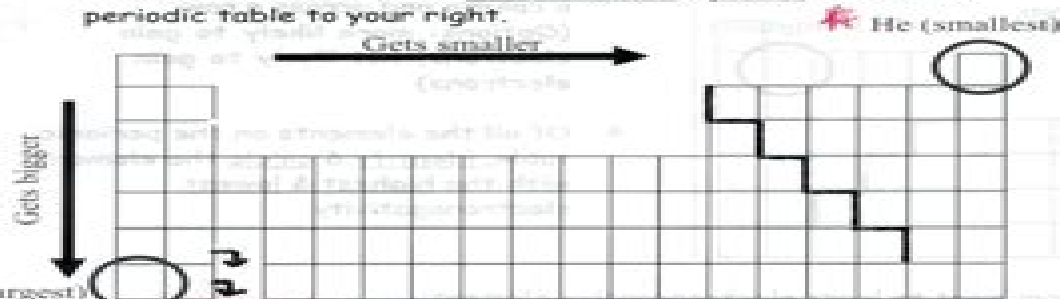


Neon
Part of the Noble family, Neon does not readily interact with other elements. With a full set of valence electrons, Neon feels that he has everything he needs to make him happy without sharing with others.

Part 2: Use your textbook to complete the following:

Atomic Radii Size:

1. Explain **atomic radius**: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together.
2. Using the information in figure 5-13 (pg 141) in your book, **write** the atomic radii size for **alkali metals & period #2** on the periodic table to your right.



3. What happens to atomic radii as you move down a column and across a row? On the arrows on the diagram, indicate the trend. (Options: gets bigger OR gets smaller)
4. Of all the elements on the periodic table, **identify & circle the largest & smallest** in size.

5. Rank the following elements from largest to smallest in atomic radii:

a) Be, Ba, Mg = **Ba, Mg, Be**

b) Li, F, B = **Li, B, F**

6. How can you explain the trend?

a. Why do they get smaller as you go across a period? Additional protons in the nucleus "pull" on electrons, bring the electron cloud closer.

Effective nuclear charge, Z_{eff} , increases the attraction of the nucleus and therefore pulls the electron cloud closer to the nucleus resulting in smaller radii.

not a good enough answer
4AP Answer

Periodic Table Pogil Answers Extension Questions

Yan Bai



Periodic Table Pogil Answers Extension Questions:

This book delves into Periodic Table Pogil Answers Extension Questions. Periodic Table Pogil Answers Extension Questions is a vital topic that needs to be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Periodic Table Pogil Answers Extension Questions, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Periodic Table Pogil Answers Extension Questions
- Chapter 2: Essential Elements of Periodic Table Pogil Answers Extension Questions
- Chapter 3: Periodic Table Pogil Answers Extension Questions in Everyday Life
- Chapter 4: Periodic Table Pogil Answers Extension Questions in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Periodic Table Pogil Answers Extension Questions. This chapter will explore what Periodic Table Pogil Answers Extension Questions is, why Periodic Table Pogil Answers Extension Questions is vital, and how to effectively learn about Periodic Table Pogil Answers Extension Questions.
3. In chapter 2, the author will delve into the foundational concepts of Periodic Table Pogil Answers Extension Questions. The second chapter will elucidate the essential principles that must be understood to grasp Periodic Table Pogil Answers Extension Questions in its entirety.
4. In chapter 3, this book will examine the practical applications of Periodic Table Pogil Answers Extension Questions in daily life. This chapter will showcase real-world examples of how Periodic Table Pogil Answers Extension Questions can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of Periodic Table Pogil Answers Extension Questions in specific contexts. This chapter will explore how Periodic Table Pogil Answers Extension Questions is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Periodic Table Pogil Answers Extension Questions. The final chapter will summarize the key points that have been discussed throughout the book.
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Table of Contents Periodic Table Pogil Answers Extension Questions

1. Understanding the eBook Periodic Table Pogil Answers Extension Questions
 - The Rise of Digital Reading Periodic Table Pogil Answers Extension Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Pogil Answers Extension Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Pogil Answers Extension Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Pogil Answers Extension Questions
 - Personalized Recommendations
 - Periodic Table Pogil Answers Extension Questions User Reviews and Ratings
 - Periodic Table Pogil Answers Extension Questions and Bestseller Lists
5. Accessing Periodic Table Pogil Answers Extension Questions Free and Paid eBooks
 - Periodic Table Pogil Answers Extension Questions Public Domain eBooks
 - Periodic Table Pogil Answers Extension Questions eBook Subscription Services
 - Periodic Table Pogil Answers Extension Questions Budget-Friendly Options
6. Navigating Periodic Table Pogil Answers Extension Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Table Pogil Answers Extension Questions Compatibility with Devices
 - Periodic Table Pogil Answers Extension Questions Enhanced eBook Features
7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Periodic Table Pogil Answers Extension Questions
 - Highlighting and Note-Taking Periodic Table Pogil Answers Extension Questions
 - Interactive Elements Periodic Table Pogil Answers Extension Questions
8. Staying Engaged with Periodic Table Pogil Answers Extension Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Pogil Answers Extension Questions
 9. Balancing eBooks and Physical Books Periodic Table Pogil Answers Extension Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Pogil Answers Extension Questions
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Periodic Table Pogil Answers Extension Questions
 - Setting Reading Goals Periodic Table Pogil Answers Extension Questions
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Periodic Table Pogil Answers Extension Questions
 - Fact-Checking eBook Content of Periodic Table Pogil Answers Extension Questions
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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