

### Unit 3: Periodic Table Review Sheet

1. Match the vocabulary term on the left with the definition on the right

|                       |                                                                                                                                                                                  |
|-----------------------|----------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| A. Periodic Law       | 1. _____ These elements are poor conductors, they are not malleable, shiny nor ductile, and have a wide range of physical properties such as color, melting point, and hardness. |
| B. Atomic Radius      | 2. _____ the energy required to remove one of an atoms electrons                                                                                                                 |
| C. Nonmetals          | 3. _____ <i>Horizontal Row</i> – the row number is the same as the principal quantum number                                                                                      |
| D. Ionization Energy  | 4. _____ atoms tend to gain, lose or share electrons in order to gain a full set of valence electrons.                                                                           |
| E. Groups or Families | 5. _____ These elements are good conductors of heat and electricity, malleable, ductile, shiny, usually gray or black and almost always solid at room temperature.               |
| F. Valence Electrons  | 6. _____ reflects an atoms ability to attract electrons in a chemical bond                                                                                                       |
| G. Periods            | 7. _____ distance from the center of the atom's nucleus to its outermost electrons                                                                                               |
| H. Electronegativity  | 8. _____ elements with similar properties are arranged in <i>vertical columns</i> . These elements have similar electron configurations                                          |
| I. Metals             | 9. _____ When elements are arranged in order of increasing atomic number, their physical and chemical properties show a periodic trend                                           |
| J. Octet Rule         | 10. _____ occupy the highest principal energy level and are largely responsible for an atoms chemical properties                                                                 |

- What information can be found in each block of the periodic table?
- Outline the history of the development of the periodic table. Include the contributions of Dobereiner, Newland, Mendeleev, and Moseley.
- Describe the properties of metalloids:
- Fill in the Blanks in the Chart Below (Hint: This is a Dobereiner Triad)

| Element | Atomic mass (amu) | Density (g/cm <sup>3</sup> ) | Melting point (°C) | Boiling point (°C) |
|---------|-------------------|------------------------------|--------------------|--------------------|
| A       | 30                | 2                            | 150                |                    |
| B       |                   | 4                            | 200                | 500                |
| C       | 60                |                              |                    | 600                |

6. Predict the ion formed by each of the following elements?

- |               |              |
|---------------|--------------|
| A. Krypton    | E. Magnesium |
| B. Silicon    | F. Chlorine  |
| C. Aluminum   | G. Cesium    |
| D. Phosphorus | H. Sulfur    |

# Periodic Table Review Sheet Answers

**Lei Huang**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name 'Lei Huang'.

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