

Periodicity

Periodicity is a regularly repeating pattern of atomic, physical and chemical properties with increasing atomic number.

> Period

The elements in a period exhibit periodicity.

We can best illustrate periodicity by looking at the elements in Periods 2 and 3.

As you move from left to right the elements go from highly electropositive metals through metalloids with giant structures to the simple molecular structure of non-metals.

Li	Be	B	C	N₂	O₂	F₂	Ne
< - <i>metals</i> - >		<i>metalloid</i>	<i>giant molecule</i>	< <i>non metals (simple molecules)</i> >			
Na	Mg	Al	Si	P₄	S₈	Cl₂	Ar
< - - - - <i>metals</i> - - - - >		<i>metalloid</i>		< - <i>non metals (simple molecules)</i> - >			

Initially one is filling the 3s orbital then the 3p orbitals.

The nuclear charge increases by one each time giving an increased pull on the electrons.

> Atomic radii

The atomic radius of an element is a measure of the size of its atoms. It is the distance from the centre of the nucleus to the boundary of the electron cloud.

Since the atom does not have a well-defined boundary, we can find the atomic radius by determining the distance between the two nuclei and dividing it by two.



Fig A. Measuring the covalent and van der Waals radii.

The diagram on the left shows two bonded atoms. The atoms are closer together than they are in the right-hand diagram (where they are just touching). On the left-hand side of the diagram, we are measuring the *covalent radius*.

The radius being measured in the diagram on the right is called the *van der Waals radius*. This is the only radius that can be determined for neon and argon, because they do not bond with other elements.

There is a third radius that is used for metals. It is called the *metallic radius* (half the distance between the inter-nuclear distance of ions).

The diagram below shows the trend in covalent radii across the second period (lithium to fluorine, or Li to F) and third period (sodium to chlorine, or Na to Cl).

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Blaine E. Hatt



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