

Chemistry I: Periodicity Worksheet

- 1) Arrange the following elements in order of increasing size: Al, Br, C, K, and Na

Discussion

- 2) Arrange the following elements in order of increasing size: Ca, Rb, P, Ge, and Sr.

Keywords

33. Circle the atom or ion in each pair that has the larger radius.

- ch. 11 or 12

- 4) Select the atom or ion in each pair that has the largest radius.

- ch. Error As

- 5) Which of the following groups of elements is arranged correctly in order of increasing ionization energy?
(Answer by circling the correct answer.)

- (d)
- $\text{Ne} < \text{C} < \text{Si} < \text{Li}$

- 6) Arrange the following atoms in order of increasing ionization energy: Li, K, C, and N.

Abstract

- 7) Arrange the following atoms in order of increasing ionization energy: Si, K, As, and Ca.

[illegible]

- 8) Compare the elements Na, Mg, O, and P.

- Background:** The purpose of this study was to determine the prevalence of and risk factors for

Periodicity Worksheet Answer Key

RJ Shavelson

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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