

Complete in the following table:

Correct shapes included  
although they're only necessary if  
specifically asked for.

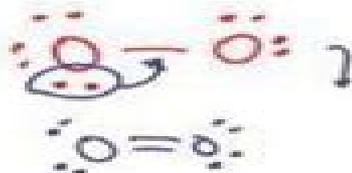
Group:	IA	IVA	VA	VIA	VIIA
Element:	H	C	N	O	F/Cl/Br/I
# valence electrons:	1	4	5	6	7
Normal # covalent bonds:	1	4	3	2	1

For the following neutral molecules, calculate the total number of valence electrons and draw the correct Lewis structure. (For neutral molecules, it's okay to assume the normal numbers of covalent bonds apply and that the central atom is the first non-H element in the formula.)

1.  $F_2$   $7 + 7 = 14$  ve total



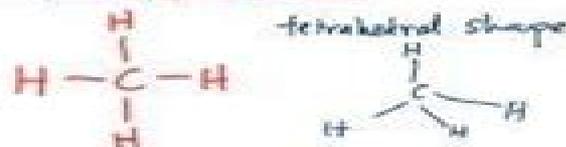
2.  $O_2$   $6 + 6 = 12$  ve total



3.  $N_2$   $5 + 5 = 10$  ve total



4.  $CH_4$   $4 + 4(1) = 8$  ve total



5.  $CH_3Cl$   $4 + 3(1) + 7 = 14$  ve total



6.  $CH_2Cl_2$   $4 + 2(1) + 2(7) = 20$  ve total



# South Pasadena Bonding And Molecular Structure Worksheet Answers

**J. K. Sirkin, Y. K. Syrkin, M. E. Diatkina**

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