

## 9

## A PERIODIC TABLE LOGIC PROBLEM

Small-Scale Experiment for text Section 6.3

**OBJECTIVES**

- Analyze properties and relationships of 26 elements.
- Identify each element's place in the periodic table based on these properties and relationships.

**INTRODUCTION**

When elements are arranged in the periodic table in order of increasing atomic number, there is a periodic repetition of their physical and chemical properties. The way properties change from left to right across a period and from top to bottom within a group are called periodic trends. Here is a summary of some periodic trends.

Atomic size can be expressed by a measurement called the atomic radius. Atomic radius is one half the distance between the nuclei of two atoms of the same element when the atoms are joined. The atomic radius generally increases from top to bottom within a group. The atomic radius generally decreases from left to right across a period.

Ionization energy is the energy required to remove an electron from an atom. This energy is measured when the atom is in the gaseous state. The first ionization energy is the energy required to remove the first electron from an atom. The second ionization energy is the energy required to remove an electron from a cation with a  $1+$  charge. The third ionization energy is the energy required to remove an electron from a cation with a  $2+$  charge. The second ionization energy for an element is always larger than the first ionization energy. The third ionization energy is always larger than the second ionization energy. In general, first ionization energy decreases from top to bottom within a group and increases from left to right across a period.

Electronegativity is the ability of an atom of an element to attract electrons when the atom is in a compound. In general, electronegativity values decrease from top to bottom within a group and increase from left to right across a period. Because noble gases do not tend to form compounds, electronegativity values for noble elements are usually omitted from data tables.

Metals tend to have low first ionization energies and low electronegativity values. Nonmetals tend to have high first ionization energies and high electronegativity values.

# Small Scale Experiment For Text Section 6 3

**Roger S. V. Pullin**



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