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web oxidation reduction redox reactions oxidation $\text{Cu} \rightarrow \text{Cu}^{2+}$ reduction $\text{MnO}_4^- \rightarrow \text{Mn}^{2+}$ when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} \rightarrow \text{MnO}_4^-$ is

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} + \text{CN}^- + 2\text{e}^- + 2\text{H}^+ + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{CN}^-$ make electrons equal convert to basic solution $4\text{CN}^- + 2\text{Au} + 2\text{CN}^- + 2\text{e}^-$ multiplied by a factor of 2 $2\text{e}^- + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{OH}^-$ the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^{-}$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 \rightarrow \text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} + 2\text{NO}$ choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO b NO_2 and $\text{C}_6\text{H}_8\text{O}_6$ c $\text{C}_6\text{H}_6\text{O}_6$ and NO_2 d $\text{C}_6\text{H}_6\text{O}_6$ and NO

2 and c 6 h 8 o 6 b no 2 and c 6 h 8 o 6 h and no 2 c h and

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