

General Chemistry
David McDermott, Peter A. Mol, John HolmanUniversity Science Books
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Map

Use the data given here to calculate the values of $\Delta G^\circ_{\text{rxn}}$ at 25 °C for the reaction described by the equation



Compound	ΔG°_f (kJ/mol)
A	+387.7
B	-603.3
C	+402.0

$$\Delta G^\circ_{\text{rxn}} = \boxed{617.6} \text{ kJ}$$

If $\Delta H^\circ_{\text{rxn}}$ and $\Delta S^\circ_{\text{rxn}}$ are both negative values, what drives the spontaneous reaction and in what direction at standard conditions?

Explanation

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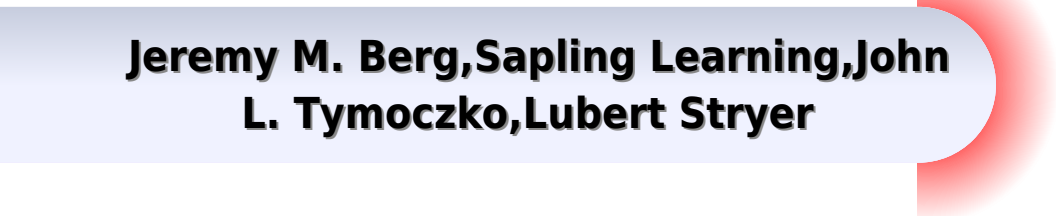
$$\Delta G^\circ_{\text{rxn}} = (1 \text{ mol}) \left(402.0 \frac{\text{kJ}}{\text{mol}} \right) - \left[(1 \text{ mol}) \left(387.7 \frac{\text{kJ}}{\text{mol}} \right) + (1 \text{ mol}) \left(-603.3 \frac{\text{kJ}}{\text{mol}} \right) \right] = 617.6 \text{ kJ}$$

Spontaneous reactions have a negative value for the Gibbs energy change. The sign of $\Delta G^\circ_{\text{rxn}}$ for the reaction as written (to the right) is positive. Thus, the reaction is spontaneous to the left.

Enthalpy or entropy alone does not determine reaction spontaneity, but we can identify the signs of ΔH and ΔS that contribute favorably to spontaneity. Exothermic reactions (negative ΔH_{rxn}) are favored. An increase in entropy (positive ΔS_{rxn}) is favorable. If a reaction has both of these, it is certainly spontaneous as written. But if a reaction has one favorable quantity and one unfavorable quantity, it may still be spontaneous as written, driven by the favorable quantity. If the reaction is spontaneous in the reverse direction, we'd consider the opposite signs of ΔH and ΔS , and so the quantity that was unfavorable in the forward direction becomes favorable (and thus the driving force) in the reverse direction.

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**Jeremy M. Berg, Sapling Learning, John
L. Tymoczko, Lubert Stryer**



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