
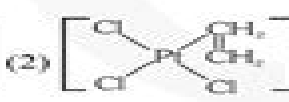




Previous Years AIPMT/NEET & AIIMS Questions

1. CuSO_4 when reacts with KCN forms CuCN , which is insoluble in water. It is soluble in excess of KCN, due to formation of the following complex :- [AIPMT 2002]
 (1) $\text{K}_2[\text{Cu}(\text{CN})_4]$ (2) $\text{K}_3[\text{Cu}(\text{CN})_4]$ (3) CuCN_2 (4) $\text{Cu}[\text{KC}u(\text{CN})_4]$
2. CN^- is a strong field ligand. This is due to the fact that : [AIPMT-2004]
 (1) It forms high spin complexes with metal species
 (2) It gives negative charge
 (3) It is a pseudohalide
 (4) It can accept electrons from metal species
3. Considering H_2O as a weak field ligand, the number of unpaired electrons in $[\text{Mn}(\text{H}_2\text{O})_6]^{2+}$ will be (At. no. of Mn = 25) :- [AIPMT-2004]
 (1) Four (2) Three (3) Five (4) Two
4. Which of the following coordination compounds would exhibit optical isomerism ? [AIPMT-2004]
 (1) tris-(ethylenediamine) cobalt (III) bromide
 (2) pentaamminenitrocobalt (III) iodide
 (3) diamminedichloroplatinum (II)
 (4) trans-dicyanobis (ethylenediamine) chromium (III) chloride
5. Among $[\text{Ni}(\text{CO})_4]$, $[\text{Ni}(\text{CN})_4]^{2-}$, $[\text{NiCl}_4]^{2-}$ species, the hybridization states at the Ni atom are, respectively (At. no. of Ni = 28) :- [AIPMT-2004]
 (1) dsp^2 , sp^3 , sp^3 (2) sp^3 , dsp^2 , dsp^2 (3) sp^3 , dsp^2 , sp^3 (4) sp^3 , sp^3 , dsp^2
6. Which of the following is considered to be an anticancer species ? [AIPMT-2004]
 (1)  (2)  (3)  (4) 
7. Which of the following does not have a metal-carbon bond ? [AIPMT-2004]
 (1) $\text{Ni}(\text{CO})_4$ (2) $\text{Al}(\text{OC}_2\text{H}_5)_3$ (3) $\text{C}_2\text{H}_5\text{MgBr}$ (4) $\text{K}[\text{Pt}(\text{C}_2\text{H}_5)\text{Cl}_3]$
8. Which one of the following is an inner orbital complex as well as diamagnetic in behaviour:- [AIPMT-2005]
 (1) $[\text{Zn}(\text{NH}_3)_6]^{2+}$ (2) $[\text{Ni}(\text{NH}_3)_6]^{2+}$
 (3) $[\text{Cr}(\text{NH}_3)_6]^{3+}$ (4) $[\text{Co}(\text{NH}_3)_6]^{3+}$
 [Atomic number : Zn = 30, Cr = 24, Co = 27, Ni = 28]

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Michal Rosen-Zvi



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