

CHEMISTRY SUBJECT TEST 3—Continued

46. Which of the following compounds would be expected to have the greatest lattice binding energy?

(A) LiNO_3
(B) LiF
(C) KI
(D) NH_4Br
(E) CsNO_3

47. The daughter nucleus formed when ^{18}F undergoes positron emission is

(A) ^{18}N
(B) ^{18}O
(C) ^{18}O
(D) ^{18}F
(E) ^{20}Ne

48. Which of the following elements would be the strongest reducing agent?

(A) Fluorine
(B) Nickel
(C) Xenon
(D) Phosphorus
(E) Sodium

49. Which of the following reactions produces a yellow precipitate?

(A) $\text{NaOH}(aq) + \text{HCl}(aq) \rightarrow \text{NaCl}(s) + \text{H}_2\text{O}$
(B) $\text{NaOH}(aq) + \text{BaCl}_2(aq) \rightarrow \text{BaOH}(s) + \text{NaCl}(aq)$
(C) $\text{Pb}(\text{NO}_3)_2(aq) + 2\text{KI}(aq) \rightarrow 2\text{KNO}_3(aq) + \text{PbI}_2(s)$
(D) $\text{CuO}(s) + \text{Mg}(s) \rightarrow \text{Cu}(s) + \text{MgO}(s)$
(E) $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$



50. In the electrochemical cell described by the cell diagram above, what reaction occurs at the anode?

(A) $\text{Zn} \rightarrow \text{Zn}^{2+} + 2e^-$
(B) $\text{Zn}^{2+} + 2e^- \rightarrow \text{Zn}$
(C) $\text{Cl}_2 + 2e^- \rightarrow 2\text{Cl}^-$
(D) $2\text{Cl}^- \rightarrow \text{Cl}_2 + 2e^-$
(E) $\text{Zn} + \text{Cl}_2 \rightarrow \text{ZnCl}_2$

51. Given the reaction $\text{A} \rightarrow \text{B} + \text{C}$, where ΔH_{rxn} is negative, what effect would increasing the temperature (at constant pressure) have on the system at equilibrium?

(A) No change
(B) Cannot be determined
(C) Shift to the right
(D) Shift to the left for $K < 1$ and to the right for $K > 1$
(E) Shift to the left

52. An unknown acid solution was presumed to be either HCl or H_2SO_4 . Which one of the following salt solutions would produce a precipitate when added to H_2SO_4 but not when added to HCl ?

(A) LiNO_3
(B) NH_4NO_3
(C) CsNO_3
(D) $\text{Ba}(\text{NO}_3)_2$
(E) AgNO_3



53. What is the equilibrium expression for the dissolution of $\text{Ca}_3(\text{PO}_4)_2$ where the above is true?

(A) $K_{\text{sp}} = [\text{Ca}^{2+}][\text{PO}_4^{3-}]^2$
(B) $K_{\text{sp}} = [\text{Ca}^{2+}]^3[\text{PO}_4^{3-}]^2$
(C) $K_{\text{sp}} = [\text{Ca}^{2+}][\text{PO}_4^{3-}]/[\text{Ca}_3(\text{PO}_4)_2]$
(D) $K_{\text{sp}} = [\text{Ca}^{2+}]^3[\text{PO}_4^{3-}]^2/[\text{Ca}_3(\text{PO}_4)_2]$
(E) $K_{\text{sp}} = [\text{Ca}^{2+}][\text{PO}_4^{3-}]^2/[\text{Ca}_3(\text{PO}_4)_2]$

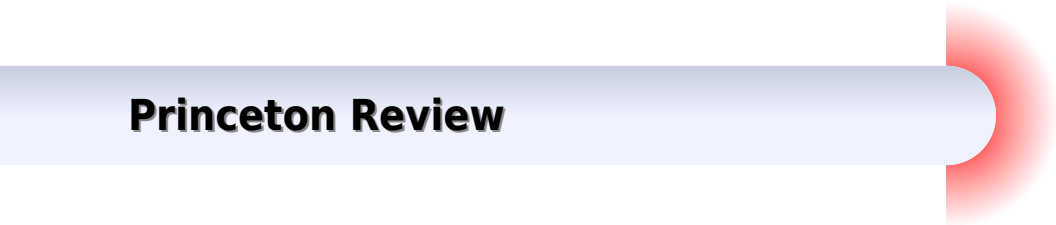
54. Which of the following represents a conjugate acid/base pair?

(A) Na^+/Cl^-
(B) HCl/H^+
(C) $\text{H}_2\text{CO}_3/\text{CO}_3^{2-}$
(D) $\text{NH}_3/\text{NH}_4^+$
(E) K^+/OH^-

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