SAT Chemistry Practice Test 15

SAT Chemistry Practice Test 1: Part C

 Twenty-five percent of element X exists as ²¹⁰X and 75 percent of it exists as ²¹⁴X. What is the atomic weight of element X in AMU?

- A. 85
- B. 211
- C. 212
- D 213
- E. 214

 A 600-milliliter container holds 2 moles of O₂(g), 3 moles of H₂(g), and 1 mole of He(g). Total pressure within the container is 760 torr. What is the partial pressure of O₅?

- A. 127 torr
- B. 253 torr
- C: 380 torr.
- D. 507 torr
- E. 760 torr

The ionic solid Fe(OH)_a is added to water and dissociates into its component ions, as shown above. The solubility product expression for the saturated solution is

- A. $K_{nn} = [Fe^{3\pi}][OHT]$
- B. $K_{ao} = [Fe^{3\pi}][3OH^*]$
- $C. K_{**} = [Fe^{3*}] [3OHT]^3$
- D. $K_{ae} = [Fe^{3+}][OHT]^3$

 $E. K_{so} = [Fe(OH),]$

4. Which of the following electron configurations represents an atom of magnesium in an excited state?

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