

Endpoint:

18.5 Solubility Equilibria and the Solubility Product Constant (p. 783-788)

Solubility Product Constant:

Molar Solubility:

In general, the solubility of an ionic compound is _____ in a solution containing a common ion than in pure water.

In general, the solubility of an ionic compound with a strongly basic or weakly basic anion _____ with increasing acidity (decreasing pH).

18.6 Precipitation (p. 789-792)

When Q is less than K_{sp} , this is an _____ solution.

When Q is equal to K_{sp} , this is a _____ solution.

When Q is greater than K_{sp} , this is a _____ solution.

Selective Precipitation:

18.7 Qualitative Chemical Analysis (p. 792-794)

Qualitative Analysis:

Quantitative Analysis:

Wet Chemistry:

18.8 Complex Ion Equilibria (p. 795-798)

Complex Ion:

Ligand:

Formation Constant:

The solubility of an ionic compound containing a metal cation that forms complex ions _____ in the presence of Lewis bases that complex with the cation.

Amphoteric:

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CL Gary



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