

THE MOLE AND AVOGADRO'S NUMBER

Name _____

One mole of a substance contains Avogadro's Number (6.02×10^{23}) of molecules.

How many molecules are in the quantities below?

1. 2.0 moles $x = 2.0 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{1.2 \times 10^{24} \text{ molecules}}$

2. 1.5 moles $x = 1.5 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{9.0 \times 10^{23} \text{ molecules}}$

3. 0.75 mole $x = 0.75 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{4.5 \times 10^{23} \text{ molecules}}$

4. 15 moles $x = 15 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{9.0 \times 10^{24} \text{ molecules}}$

5. 0.35 mole $x = 0.35 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{2.1 \times 10^{23} \text{ molecules}}$

How many moles are in the number of molecules below?

1. 6.02×10^{23} $x \text{ moles} = 6.02 \times 10^{23} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{1.0 \text{ mol}}$

2. 1.204×10^{24} $x \text{ moles} = 1.204 \times 10^{24} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{2.0 \text{ mol}}$

3. 1.5×10^{20} $x \text{ moles} = 1.5 \times 10^{20} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{2.5 \times 10^{-4} \text{ mol}}$ (0.00025 mol)

4. 3.4×10^{26} $x \text{ moles} = 3.4 \times 10^{26} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{565 \text{ mol}}$

5. 7.5×10^{19} $x \text{ moles} = 7.5 \times 10^{19} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{1.2 \times 10^{-4} \text{ mol}}$

Shapes Of Molecules Chemistry If8766 Answers

Theodore W. Gray



Shapes Of Molecules Chemistry If8766 Answers:

The Shape and Structure of Molecules Charles Alfred Coulson, 1982 Molecular Geometry Alison Rodger, Mark Rodger, 2014-05-16 Molecular Geometry discusses topics relevant to the arrangement of atoms. The book is comprised of seven chapters that tackle several areas of molecular geometry. Chapter 1 reviews the definition and determination of molecular geometry while Chapter 2 discusses the unified view of stereochemistry and stereochemical changes. Chapter 3 covers the geometry of molecules of second row atoms and Chapter 4 deals with the main group elements beyond the second row. The book also talks about the complexes of transition metals and f block elements and then covers the organometallic compounds and transition metal clusters. The last chapter tackles the consequences of small local variations in geometry. The text will be of great use to chemists who primarily deal with the properties of molecules and atoms. **Molecules and Models** Arne Haaland, 2008-03-06 This book provides a systematic description of the molecular structures and bonding in simple compounds of the main group elements with particular emphasis on bond distances, bond energies and coordination geometries. The description includes the structures of hydrogen halogen and methyl derivatives of the elements in each group; some of these molecules are ionic, some polar covalent. The survey of molecules whose structures conform to well established trends is followed by representative examples of molecules that do not conform. We also describe electron donor-acceptor and hydrogen bonded complexes. Chemists use models to systematize our knowledge, to memorize information and to predict the structures of compounds that have not yet been studied. The book provides a lucid discussion of a number of models such as the Lewis electron pair bond and the VSEPR models, the spherical and polarizable ion models and molecular orbital calculations and it outlines the successes and failures of each. **The Structure of Small Molecules** W. J. Orville-Thomas, 1966 **Shape in Chemistry** Paul G. Mezey, 1993 **Molecules** Dale Atkins, 1987-08-15 Explains how diagrams are used to represent chemical bonds and describes the structure and characteristics of molecules encountered in everyday life. **Molecules** Theodore W. Gray, 2014 In his highly anticipated sequel to *The Elements*, Theodore Gray demonstrates how the elements of the periodic table combine to form the molecules that make up our world. Everything physical is made up of the elements and the infinite variety of molecules they form when they combine with each other. In *Molecules*, Theodore Gray takes the next step in the grand story that began with the periodic table in his best-selling book *The Elements: A Visual Exploration of Every Known Atom in the Universe*. Here he explores through fascinating stories and trademark stunning photography the most interesting, essential, useful and beautiful of the millions of chemical structures that make up every material in the world. Gray begins with an explanation of how atoms bond to form molecules and compounds as well as the difference between organic and inorganic chemistry. He then goes on to explore the vast array of materials molecules can create, including soaps and solvents, goops and oils, rocks and ores, ropes and fibers, painkillers and dangerous drugs, sweeteners, perfumes and stink bombs, colors and pigments and controversial compounds including asbestos.

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Molecules Peter W. Atkins, 1987 *The Determination of Molecular Structure* Peter Jaffrey Wheatley, 1962 **Atoms and Molecules**, 1995 **The Architecture of Molecules** Linus Pauling, Roger Hayward, 1964 An introduction to the subject of how atoms are arranged and interconnected in molecules and crystals and to the way in which the geometry of this organization accounts for some of the properties of substances The Author and Roger Hayward the illustrator emphasize the significance of molecular structure to life

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