

ii) Mg^{2+}

iii) N^{3-}

Q-8: What information did Kossel provide about chemical bonding?

Q-9: State with reasons

a) Which is more acidic - anhydrous HCl or aqueous HCl ?

b) Which is more polar - CO_2 or N_2O ?

Q-10: AlCl_3 is covalent when anhydrous. Determine whether it would remain covalent or change to an ionic state in an aqueous solution using the information provided below. Ionisation energy for $\text{AlCl}_3 = 5137 \text{ kJ/mol}$, $\Delta H_{\text{hydration}}$ for $\text{Al}^{3+} = -4665 \text{ kJ/mol}$, $\Delta H_{\text{hydration}}$ for $\text{Cl}^- = -381 \text{ kJ/mol}$.

Q-11: Draw the molecular structures of OsF_6 and XeF_4 using the VSEPR theory, highlighting the location of lone pairs of electrons and the hybridisation of the central atom.

Q-12: Determine the HCl 's ionic character percentage. Given that the HCl 's bond length is 1.275 and the observed dipole moment is 1.83 D .

Q-13: Describe the change in hybridisation, if any.

a) Of Al atom in the reaction $\text{AlCl}_3 + \text{Cl}^- \rightarrow \text{AlCl}_4^-$?

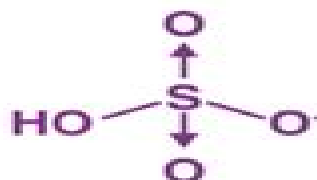
b) Of B and N atoms in the reaction $\text{BF}_3 + \text{NH}_3 \rightarrow \text{F}_3\text{B}-\text{NH}_3$?

Q-14: Describe how Lewis' concept differs from VB theory.

Q-15: Out of sp -hybrid orbital and p orbital, which has greater directional character and why?

Q-16: When we state that the lattice energy of NaCl is 788 kJ/mol . What does it mean?

Q-17: Calculate the formal charge of S in the following structure.



Q-18: Give two resonating structures of ozone which satisfy the octet rule. Also, give the probable hybrid structure.

Q-19: i) Why does MgCl_2 have a much higher solubility than MgF_2 ?

ii) Why is the melting point of MgO (2800°C) much higher than that of BaO (1920°C)?

Q-20: How will you differentiate between the lone pair and a bond pair of electrons?

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