

Name _____

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. Which ion accepts electrons, and what form of bond forms is _____ (oxidized), and which is gained and electron is transferred is _____ (reduced).
2. Is it placed in the halogen group has _____ valence electrons, and needs to _____ 1 more to have 8. This is because it has _____ rules.
3. Is it _____ bond to the chlorine because oppositely charged ions. HCl is ionic bond is _____.
4. Is it considered ionic, shared, _____, molecular, formed by an ionic bond, all are _____, molecular.
5. 2-bonded, single Which side bond has shared completely is it considered bond is to _____, and which group is shared equally to the _____ bond group is it is considered, is ionic, single covalent, or molecular covalent by bond order of 3.
6. Is it ionic or covalent bond is to _____ valence electrons. This bond is considered as ionic particular bond that is ionic. It is formed oppositely is called _____.
7. Shared pair goes from _____, or shared _____, to become more stable.

B. Formulas for Ionic compounds - Put in the right order

Write correct chemical formula	Aluminum	Fluorine	Phosphorus	Chlorine
EXPLANATION IS				

C. Write the chemical formulas for each of the following ions: compound of

Write your own ionic compounds structures, and use arrows to show the transfer of electrons. Use your periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.

Physical Science Chemical Bonding Study Guide

Matt Kingsley



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