# TABLE 3: THE PERIODIC TABLE OF ELEMENTS TABEL 3: DIE PERIODIEKE TABEL VAN ELEMENTE

|     | 0               |      | 2<br>(II)       |     | 3   |     | 4               | 5               | 6  | 1                | 8                 | 9                  | 10                | 11                 | 12                | 13<br>(III)          | 14<br>(IV)        | 15<br>(V)         | 16<br>(VI)         | 17<br>(VII)       | 18<br>(VIII)    |
|-----|-----------------|------|-----------------|-----|---|-----|-----------------|-----------------|--|------------------|-------------------|--------------------|-------------------|--------------------|-------------------|----------------------|-------------------|-------------------|--------------------|-------------------|-----------------|
| ,   | 1<br>H          |      | .000            |     |   |     |                 | KEY/SŁ          | EUTEL                                    | ,                | Afoom             |                    |                   |                    |                   | VV                   | W.7.6             | 1*1               | 17.00              | 11.11             | He<br>4         |
| 9.  | J<br>Li<br>7    | 1,5  | 4<br>Be         |     |   |     |                 |                 | Electronegativity<br>Elektronegatiwiteif |                  |                   | 29 Cu<br>53,5 + Si |                   |                    |                   | 5<br>3 B<br>11       | % C<br>12         | 7<br>2 N<br>14    | 8<br>2 O<br>16     | 9<br>9 F<br>19    | 10<br>Ne<br>20  |
| ô   | 11<br>Na<br>23  | 17.5 | 12<br>Mg<br>24  |     | Approximate relative atomic mass Benaderde relatiewe atoommassa |     |                 |                 |  |                  |                   |                    |                   |                    | 13<br>2 A€<br>27  | 14<br>Si<br>28       | 15<br>F P<br>31   | 16<br>% S<br>32   | 17<br>2 C€<br>35,5 | 18<br>Ar<br>40    |                 |
| 80  | 19<br>K<br>39   | 0.7  | 20<br>Ca<br>40  | 1,5 | 21<br>Sc<br>45  | 1,6 | 22<br>Ti<br>48  | 23<br>≤ V<br>51 | 24<br>5 Cr<br>52                         | 25<br>S Mn<br>56 | 26<br>3 Fe<br>56  | 27<br>5 Co<br>59   | 28<br>S Ni<br>59  | 29<br>5 Cu<br>63,5 | 30<br>\$ Zn<br>65 | 31<br><b>5 Ga</b> 70 | 32<br>3 Ge<br>73  | 33<br>2 As<br>75  | 34<br>3 Se<br>79   | 35<br>23 Br<br>80 | 36<br>Kr<br>84  |
| 0,0 | 37<br>Rb<br>86  | 0.   | 38<br>Sr<br>88  | 2   | 39<br>Y<br>89   | *   | 40<br>Zr<br>91  | 41<br>Nb<br>92  | 42<br>\$ Mo<br>96                        | 43<br>□ Tc       | 44<br>2 Ru<br>101 | 45<br>⊋ Rh<br>103  | 46<br>□ Pd<br>106 | 47<br>\$ Ag<br>108 | 48<br>□ Cd<br>112 | 49<br>□ In<br>115    | 50<br>2 Sn<br>119 | 51<br>Sb<br>122   | 52<br>☐ Te<br>128  | 53<br>% I<br>127  | 54<br>Xe<br>131 |
| 5   | 55<br>Cs<br>133 | 6.0  | 56<br>Ba<br>137 |     | 57<br>La<br>139   | 1.6 | 72<br>Hf<br>179 | 73<br>Ta<br>181 | 74<br>W<br>184                           | 75<br>Re<br>186  | 76<br>Os<br>190   | 77<br>Ir<br>192    | 78<br>Pt<br>195   | 79<br>Au<br>197    | 80<br>Hg<br>201   | 81<br>□ T€<br>204    | 82<br>2 Pb<br>207 | 83<br>2 Bi<br>209 | 84<br>2 Po         | 85<br>2 At        | 86<br>Rn        |
| 3   | 87<br>Fr        | 6.0  | 88<br>Ra<br>226 |     | Ac  |     |                 | 58<br>Ce<br>140 | 59<br>Pr<br>141                          | 60<br>Nd<br>144  | 61<br>Pm          | 62<br>Sm<br>150    | 63<br>Eu<br>152   | 64<br>Gd<br>157    | 65<br>Tb<br>159   | 65<br>Dy<br>163      | 67<br>Ho<br>165   | 68<br>Er<br>167   | 69<br>Tm<br>169    | 70<br>Yb<br>173   | 71<br>Lu<br>175 |
|     |                 |      |                 |     |   |     |                 | 90<br>Th<br>232 | 91<br>Pa                                 | 92<br>U<br>238   | 93<br>Np          | 94<br>Pu           | 95<br>Am          | 96<br>Cm           | 97<br>Bk          | 98<br>Cf             | es<br>Es          | Fm                | Md<br>Md           | No<br>No          | 103<br>Lr       |

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