

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

1. Physical Science Basics

- What tool(s) would you want to use to find the:
 - Mass of an object? triple beam balance Basic SI Unit? kg
 - Volume of an object? ruler or water displacement method Basic SI Unit? cm³ or ml
 - Length of an object? Ruler Basic SI Unit? meter
 - Density of an object? graduated cylinder Basic SI Unit? ml
- What is the formula for calculating density? mass/volume
 - What is the density of a cube that is 25 g and has a volume of 3 cm³? 8g/cm³
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has.
solid H₂O less dense than liquid H₂O

2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1802	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell (<u>energy level</u>) to another.	Bohr Model
Ernest Rutherford	1909	1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
 - An element's properties can be predicted from its location in the periodic table.
 - The atomic number represents the number of protons.
 - The atomic mass represents the number of protons and neutrons.

Physical Science 2nd Semester Exam Review

United States Air Force Academy

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the United States Air Force Academy text.

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Essentials of Physical Chemistry Don Shillady, 2011-07-27 At a time when U S high school students are producing low scores in mathematics and science on international examinations a thorough grounding in physical chemistry should not be considered optional for science undergraduates Based on the author s thirty years of teaching *Essentials of Physical Chemistry* merges coverage of calculus with chemistry and molecular physics in a friendly yet thorough manner Reflecting the latest ACS guidelines the book can be used as a one or two semester course and includes special topics suitable for senior projects The book begins with a math and physics review to ensure all students start on the same level and then discusses the basics of thermodynamics and kinetics with mathematics tuned to a level that stretches students abilities It then provides material for an optional second semester course that shows students how to apply their enhanced mathematical skills in a brief historical development of the quantum mechanics of molecules Emphasizing spectroscopy the text is built on a foundation of quantum chemistry and more mathematical detail and examples It contains sample classroom tested exams to gauge how well students know how to use relevant formulas and to display successful understanding of key concepts Coupling the development of mathematical skills with chemistry concepts encourages students to learn mathematical derivations Mini biographies of famous scientists make the presentation more interesting from a people point of view Stating the basic concepts of quantum chemistry in terms of analogies provides a pedagogically useful technique Covering key topics such as the critical point of a van der Waals gas the Michaelis Menten equation and the entropy of mixing this classroom tested text highlights applications across the range of chemistry forensic science pre medical science and chemical engineering In a presentation of fundamental topics held together by clearly established mathematical models the book supplies a quantitative discussion of the merged science of physical chemistry

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A Course on Set Theory Ernest Schimmerling, 2011-07-28 Set theory is the mathematics of infinity and part of the core curriculum for mathematics majors This book blends theory and connections with other parts of mathematics so that readers can understand the place of set theory within the wider context Beginning with the theoretical fundamentals the author proceeds to illustrate applications to topology analysis and combinatorics as well as to pure set theory Concepts such as Boolean algebras trees games dense linear orderings ideals filters and club and stationary sets are also developed Pitched specifically at undergraduate students the approach is neither esoteric nor encyclopedic The author an experienced instructor includes motivating examples and over 100 exercises designed for homework assignments reviews and exams It is appropriate for undergraduates as a course textbook or for self study Graduate students and researchers will also find it useful as a refresher or to solidify their understanding of basic set theory

The Hidden Curriculum—Faculty-Made Tests in Science Sheila Tobias, Jacqueline Raphael, 2013-06-29 This resource manual for college level science instructors reevaluates the role of testing in their curricula and describes innovative techniques pioneered by other teachers part I examines the effects of the following on lower division courses changes in exam content format and environment revisions in grading practices student response colleague reaction the sharing of new practices with other interested professionals and more The book includes a comprehensive introduction faculty composed narratives commentaries by well known science educators and a visual index to 100 more refined innovations

Mammography and Breast Imaging PREP: Program Review and Exam Prep Olive Peart, 2012-01-18 A comprehensive review for the mammography registry examination from an experienced educator and clinician who knows exactly what it takes to pass Includes new coverage of the latest digital imaging technologies Written by an instructor and mammography specialist at Stamford Hospital Concise narrative text helps you to focus on essential concepts Practice questions with answers referenced to the text allow you to gauge your comprehension of important material Learning aids such as objectives and glossaries at the beginning of each chapter streamline the learning process Numerous radiographs teach you

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