

Periodic Trends Worksheet

Directions: Use the table to determine the following properties:

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

K, Na, O, S, Cl

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

K, Na, S, O, Cl

3. Which atom has the highest first ionization energy than sodium?

Chlorine (because it has the most electronegativity and is closest to sodium)

4. Which two elements in the same family generally have similar properties?

There have been two atoms that share a same position with potassium

5. Arrange **halogens** the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (increases to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms of the periodic table? Write numbers, this period?

Periods 1 through 7 showing elements arranged in the periodic table

7. Which period in the periodic table shows atoms arranged in the periodic table? Which period shows elements?

Periods 1 through 7 showing elements arranged in the periodic table

8. Write the atoms in each group that have the highest values:

a.	Li	Be
b.	Na	Mg
c.	K	Ca
d.	Rb	Sr
e.	Cs	Ba
f.	Fr	Ra

a.	Na	Ne
b.	Ar	Cl
c.	Kr	Br
d.	Xe	I
e.	Rn	At

9. Write the atoms in each group that have the highest electronegativity:

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Denise Eby



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