3 Assumptions of The Kinetic Theory - Section 1

- 1. All matter is made of small particles
- Those particles are in constant, random motion
- 3. Those particles collide with each other and the walls of their containers
- 4. Thermal Energy total energy of a material's particles
- 4a. Kinetic Energy energy a moving object has due to its motion
- 4b. Potential Energy stored energy an object has due to its position
- 5. Average Kinetic Energy temperature; how fast the particles of a substance are moving
- Solid State tightly packed particles
- 7. Melting Point point where particles break their ordered arrangement from the solid state
- 7a. Heat of Fusion energy required to reach the melting point
- 8. Liquids Flow viscosity; the resistance of a liquid to flow affected by temp. & molecular structure
- Gas State a liquid's particles have enough energy to escape the attractive forces of other particles in a liquid
- 9a. Evaporation -
- 10. Boiling Point point where particles break free from each other from the liquid state
- 10a. Heat of Vaporization energy required to reach the boiling point
- 11. Diffusion the tendency of particles to spread evenly throughout the container/space
- Plasma State state of matter consisting of high-temperature gas with balanced positively and negatively charged particles.
- Thermal Expansion size increase of a material when its temperature increases
- 14. Water's Strange Expansion water expands as it cools
- Melting Amorphous Solids don't have a definite melting point like glass or plastic

Section 2 - Properties of Fluids

- 1. FALSE, gas
- 2. FALSE, upward
- 3. FALSE, less than
- 4 TRUE
- S.FALSE, Pascal
- 6. FALSE, decreases
- 7. TRUE
- 8. According to Pascal, pressure applied to a fluid is transmitted through the fluid. Not the same for gas
- According to Archimedes...The weight of the water is 2/3 the weight of the wood because 2/3 of the block is what was displaced
- 10. The downward force decreases on top, net lifting force on the wing increases. True b/c Bernoulli's

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