

4. a.
$$\frac{10\text{g Al} \mid 1\text{mol Al} \mid 3\text{mol Cu} \mid 63.55\text{g Cu}}{26.98\text{g Al} \mid 2\text{mol Al} \mid 1\text{mol Cu}} = 35.33\text{g Cu}$$

$$\frac{20\text{g CuCl}_2 \mid 1\text{mol CuCl}_2 \mid 3\text{mol Cu} \mid 63.55\text{g Cu}}{134.45\text{g CuCl}_2 \mid 3\text{mol CuCl}_2 \mid 1\text{mol Cu}} = \boxed{9.45\text{g Cu}}$$

b.
$$\frac{30\text{g Al} \mid 1\text{mol Al} \mid 2\text{mol AlCl}_3 \mid 133.33\text{g AlCl}_3}{26.98\text{g Al} \mid 2\text{mol Al} \mid 1\text{mol AlCl}_3} = 148.25\text{g AlCl}_3$$

$$\frac{100\text{g CuCl}_2 \mid 1\text{mol CuCl}_2 \mid 2\text{mol AlCl}_3 \mid 133.33\text{g AlCl}_3}{134.45\text{g CuCl}_2 \mid 3\text{mol CuCl}_2 \mid 1\text{mol AlCl}_3} = \boxed{66.11\text{g AlCl}_3}$$

c.
$$\frac{5\text{g Al} \mid 1\text{mol Al} \mid 3\text{mol Cu} \mid 63.55\text{g Cu}}{26.98\text{g Al} \mid 2\text{mol Al} \mid 1\text{mol Cu}} = 17.67\text{g Cu}$$

~~10g / 17.67g x 100 = 56.6%~~ $10\text{g} / 17.67\text{g} \times 100 = \boxed{56.6\%}$

d.
$$\frac{10\text{g CuCl}_2 \mid 1\text{mol CuCl}_2 \mid 2\text{mol AlCl}_3 \mid 133.33\text{g AlCl}_3}{134.45\text{g CuCl}_2 \mid 3\text{mol CuCl}_2 \mid 1\text{mol AlCl}_3} = 6.61\text{g AlCl}_3$$

~~5g / 6.61g x 100 = 75.6%~~ $5\text{g} / 6.61\text{g} \times 100 = \boxed{75.6\%}$

- 5.
- a. TRUE
 - b. FALSE
 - c. $\text{AgNO}_3 \Rightarrow 169.91\text{g/mol}$
 - d. 6.02×10^{23} particles = 1 mole of particles

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