

STOICHIOMETRY NOTES WITH LIMITING REACTANT AND PERCENT YIELD

STOICHIOMETRY

Stoichiometry is the relationship between the masses of the reactants and products of chemical reactions.

- Stoichiometry is possible because of the Law of Conservation of Mass
- The law of conservation of mass states that the amount of matter before a reaction occurs is equal to the amount of matter present after the reaction occurs.
- Thanks to this idea, we can build ratios between reactants and products that allow us to calculate amounts of each

Using Ratios

Each coefficient in the balanced equation can be used to define a ratio between the number of moles of any product or the amount of mass for the new substance formed and moles of a reactant.

Consider the reaction:
 $2\text{LiCl} + \text{MgSO}_4 \rightarrow \text{Li}_2\text{SO}_4 + \text{MgCl}_2$

We can write these ratios:

moles LiCl	$\frac{2\text{LiCl}}{1\text{MgSO}_4}$	$\frac{1\text{MgSO}_4}{2\text{LiCl}}$
moles LiCl	$\frac{2\text{LiCl}}{1\text{Li}_2\text{SO}_4}$	$\frac{1\text{Li}_2\text{SO}_4}{2\text{LiCl}}$
moles LiCl	$\frac{2\text{LiCl}}{1\text{MgCl}_2}$	$\frac{1\text{MgCl}_2}{2\text{LiCl}}$
moles MgSO ₄	$\frac{1\text{MgSO}_4}{2\text{LiCl}}$	$\frac{2\text{LiCl}}{1\text{MgSO}_4}$
moles MgSO ₄	$\frac{1\text{MgSO}_4}{1\text{Li}_2\text{SO}_4}$	$\frac{1\text{Li}_2\text{SO}_4}{1\text{MgSO}_4}$
moles MgSO ₄	$\frac{1\text{MgSO}_4}{1\text{MgCl}_2}$	$\frac{1\text{MgCl}_2}{1\text{MgSO}_4}$

Example: Finding the Limiting Reactant in the reaction:
 $2\text{LiCl}(aq) + \text{MgSO}_4(aq) \rightarrow \text{Li}_2\text{SO}_4(aq) + \text{MgCl}_2(aq)$

4.37 grams of MgSO_4 react with 2.44 grams of LiCl . What is the limiting reactant and what mass of MgCl_2 can be formed?

How many moles of LiCl react? $\frac{2.44\text{g}}{42.39\text{g/mol}} = 0.0576\text{mol}$

How many moles of MgSO_4 react? $\frac{4.37\text{g}}{120.36\text{g/mol}} = 0.0363\text{mol}$

How many moles of MgCl_2 will form from LiCl ? $0.0576\text{mol LiCl} \times \frac{1\text{MgCl}_2}{2\text{LiCl}} = 0.0288\text{mol MgCl}_2$

How many moles of MgCl_2 will form from MgSO_4 ? $0.0363\text{mol MgSO}_4 \times \frac{1\text{MgCl}_2}{1\text{MgSO}_4} = 0.0363\text{mol MgCl}_2$

0.0288 grams of MgCl_2 will form and MgSO_4 is the limiting reactant.

Using Mole Ratios

From the last page, we looked at the reaction:
 $2\text{LiCl} + \text{MgSO}_4 \rightarrow \text{Li}_2\text{SO}_4 + \text{MgCl}_2$

We also defined ratios of moles of each substance using the coefficients of the balanced equation.

How do we use the ratios we defined?

How many moles of LiCl can be produced from 4 moles of MgSO_4 reactant? $4\text{mol MgSO}_4 \times \frac{2\text{LiCl}}{1\text{MgSO}_4} = 8\text{mol LiCl}$

We showed the ratio of moles of LiCl to moles of MgSO_4 multiplying it by the given number of moles of MgSO_4 . By using this ratio and the amount we are given, we determine the ratio makes sense and we get the answer of 8 LiCl .

$4\text{mol MgSO}_4 \times \frac{2\text{LiCl}}{1\text{MgSO}_4} = 8\text{mol LiCl}$

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