

Solutions + Solubility + What is concentration and how to measure it + Potential Energy Diagram

Solute - substance being dissolved

Solvent - what the solute is dissolved in

Solution - homogeneous mixture of solute and solvent

Aqueous solution - solution where water is the solvent

Solubility - amount of solute that can be dissolved in a solvent at a given temperature

Increase pressure / decrease temperature = gas solubility increasing factors

Particles are closer / particles slow down

Increase temperature / pressure doesn't matter = solid solubility increasing factors

Particles dissolve faster

Saturated Solution - solution that has the maximum amount of solute that it can dissolve at a given temperature

Unsaturated Solution - solution that has less maximum amount of solute that it can dissolve at a given temperature

Supersaturated Solution - solution that has more than maximum amount of solute that it can dissolve at a given temperature

Solution Equilibration - rate of dissociation = rate of crystallization

Crystallization - process of formation of a crystal precipitate from dissolved solute in a solution

Dissociation - process of breaking down to dissolve

Concentration - solution that has more solute in given amount of solvent

Dilute - solution that has less solute in given amount of solvent

Molarity - mol of solute / liters of solution

PPM - mass of solute / mass of solution (10^6)

Boiling point elevation - solutions that have a higher boiling point than pure solvent

Freezing point depression - solutions that have a lower freezing point than solvent

More solute = higher boiling point = lower freezing point

Thermochemistry - study of heat absorbed or released in a reaction

System - what is being investigated (the smaller one like the solute)

Surrounding - what is being measured directly (the bigger one like the solvent)

Entropy = randomness

BARF (break absorb release form) (talking about energy here)

Enthalpy (H) = overall energy of the reaction

Study Guide Solubility And Concentration

Scott Wolfson



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a group of 30 professional calligraphers were asked to define calligraphy as a profession and this is the net result calligraphy is the art of forming beautiful symbols by hand and arranging them well it s a set of skills and techniques for positioning and inscribing words so they show integrity harmony some sort of ancestry rhythm

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