## Chapter Test B

#### **Teacher Notes and Answers**

#### 8 Chemical Equations and

#### Reactions

#### TEST B

- 1. 6
- 2 a
- 3 h.
- 4 4
- 4 0
- S. B
- 7.4
- 8 8
- 9. operfficient
- 10. Isosaid
- reverable reaction.
- 12.  $A + X \rightarrow AX$
- 13.  $AX \rightarrow A + X$
- 14.  $A + BX \rightarrow AX + B$
- 15.  $AX + BY \rightarrow AY + BX$
- 16. metal hydroxides
- 17. combustion reaction.
- 13. electrolysis
- 19. activity series
- 20, chemical equation
- 21. conservation of mass
- 22. gazeous
- 23. heated
- 24. A word equation and a formula equation are both qualitative. Although they both show the reactants and products of a chemical reaction, they give no information about amounts of reactants or products. A chemical equation, however, does show the relative amounts of the reactants and products in a chemical reaction.

- 25. Three indications of a chemical reaction are the radiation of heat and light, the production of a gas, and the formation of a precipitate.
- 26. zinc salfide + corygen → zinc oxide + sulfur dioxide
- sodium iodide + chlorine sodium chloride + iodine
- 28.  $Mg + O_2 \rightarrow MgO$
- 29. CaCO<sub>3</sub> → CaO + CO<sub>2</sub>
- 30. decomposition reaction
- 31. single-displacement reaction
- 32. synthesis reaction.
- double-displacement reaction.
- 34. 2AgNO<sub>5</sub> + CuCl<sub>2</sub> → 2AgCl + Cu(NO<sub>5</sub>)<sub>2</sub>
- 35. 2PbO<sub>2</sub> → 2PbO + O<sub>2</sub>
- 36. Zn(OH)<sub>2</sub> + 2CH<sub>3</sub>COOH →

- 37.  $Mg(s) + 2H_2O(g) \rightarrow Mg(OH)_3(\alpha q) + H_2(g)$
- 38, no reaction
- Cb(g) 1 MgBrs(aq) → MgCb(aq) + Brs(l)
- Zn(z) + 2HCl(aq) → ZnClz(aq) + Hz(g)
- 41 mo reaction
- Né(s) + CuClu(aq) → NéClu(aq) + Cu(s)
- 43, no reaction
- Mg(s) + Co(NO<sub>3</sub>)s(aq) → Mg(NO<sub>3</sub>)s(aq) + Co(s)

## **Teaching Resources Chapter 36 Answer Key**

**Kieran Walsh** 

## **Teaching Resources Chapter 36 Answer Key:**

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