

## The Activity Series

Use with Chapter 9,  
Section 9.2

1. For each of the following pairs of elements, underline the one that would replace the other element in a compound.

a. calcium, tin

e. iron, copper

b. bromine, fluorine

f. iodine, cobalt

c. aluminum, potassium

g. silver, lead

d. zinc, sodium

2. For each of the following reactions, use the activity series to determine whether the reaction would take place or not. If no reaction takes place, write NR in the blank. If a reaction does take place, write the formulas for the products of the reaction. (Hint: If an active metal replaces the hydrogen in water, the hydroxide of the active metal forms.)

a.  $\text{Li(s)} + \text{Fe(NO}_3)_3\text{(aq)} \rightarrow \text{LiNO}_3 + \text{Fe}$ b.  $\text{Au(s)} + \text{HCl(aq)} \rightarrow \text{No Rxn.}$ c.  $\text{Cl}_2\text{(g)} + \text{KBr(aq)} \rightarrow \text{KCl} + \text{Br}_2$ d.  $\text{Cu(s)} + \text{Al(NO}_3)_3\text{(aq)} \rightarrow \text{No Rxn.}$ e.  $\text{Ag(s)} + \text{HBr(aq)} \rightarrow \text{No Rxn.}$ f.  $\text{Ni(s)} + \text{SnCl}_2\text{(aq)} \rightarrow \text{NiCl}_2 + \text{Sn}$ 

3. Magnesium metal can be used to remove tarnish from silver items. Silver tarnish is the corrosion that occurs when silver metal reacts with substances in the environment, especially those containing sulfur. Why would magnesium remove tarnish from silver?

Mg more reactive, would replace silver bonded to sulfur (environment) leaving silver.

4. Use the activity series for metals to explain why copper metal is used in plumbing where the water might contain compounds of many different metals.

Cu not very reactive - won't react w/ metals in water.

5. The last four metals in the activity series of metals are commonly referred to as the "coinage metals." Why would these metals be chosen over more active metals for use in coins? Why do you think some more active metals, such as zinc or nickel, are sometimes used in coins?

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# Teaching Transparency 33 Answers

**Peter Daly, Kristen Reid, Patrick  
Buckley, Elaine Doyle**



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## **Table of Contents Teaching Transparency 33 Answers**

1. Understanding the eBook Teaching Transparency 33 Answers
  - The Rise of Digital Reading Teaching Transparency 33 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Teaching Transparency 33 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Teaching Transparency 33 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Teaching Transparency 33 Answers
  - Personalized Recommendations
  - Teaching Transparency 33 Answers User Reviews and Ratings

- Teaching Transparency 33 Answers and Bestseller Lists
- 5. Accessing Teaching Transparency 33 Answers Free and Paid eBooks
  - Teaching Transparency 33 Answers Public Domain eBooks
  - Teaching Transparency 33 Answers eBook Subscription Services
  - Teaching Transparency 33 Answers Budget-Friendly Options
- 6. Navigating Teaching Transparency 33 Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Teaching Transparency 33 Answers Compatibility with Devices
  - Teaching Transparency 33 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Teaching Transparency 33 Answers
  - Highlighting and Note-Taking Teaching Transparency 33 Answers
  - Interactive Elements Teaching Transparency 33 Answers
- 8. Staying Engaged with Teaching Transparency 33 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Teaching Transparency 33 Answers
- 9. Balancing eBooks and Physical Books Teaching Transparency 33 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Teaching Transparency 33 Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Teaching Transparency 33 Answers
  - Setting Reading Goals Teaching Transparency 33 Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Teaching Transparency 33 Answers
  - Fact-Checking eBook Content of Teaching Transparency 33 Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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