

# Visualizing the Conservation of Mass

Use with Chapter 3,  
Section 3.2

1. How many potassium atoms are in the reactants of the chemical reaction shown? How many are in the products? \_\_\_\_\_
2. How many oxygen atoms are in the reactants of the chemical reaction shown? How many are in the products? \_\_\_\_\_
3. How many hydrogen atoms are in the reactants of the chemical reaction shown? How many are in the products? \_\_\_\_\_
4. Assume that the chemical reaction shown started with having a total of 17 g of potassium and water. How much potassium hydroxide and hydrogen gas will be produced by the chemical reaction? Show your work.  
\_\_\_\_\_  
\_\_\_\_\_
5. Assume that the chemical reaction shown started with having 8 grams of potassium and 8 molecules of water. How many molecules of potassium hydroxide will be produced by the chemical reaction? How many hydrogen atoms will result?  
\_\_\_\_\_  
\_\_\_\_\_
6. Assume that you are working with a chemical reaction that decomposes water into hydrogen and oxygen. You begin with 36 grams of water and end with 12 grams of oxygen. If all of the water decomposes, how many grams of hydrogen gas will result?  
\_\_\_\_\_  
\_\_\_\_\_  
\_\_\_\_\_
7. Assume that you are working with a chemical reaction that synthesizes salt from sodium and chlorine. You begin with 100 g of chlorine. You synthesize 100 g of salt. If all of the reactants were used up, how many grams of sodium did you begin with?  
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\_\_\_\_\_  
\_\_\_\_\_

# Teaching Transparency Worksheet Ionic Bonds Chapter 7 Answers

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Spaulding**



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