

Name: _____

Worksheet – ASSIGNING OXIDATION NUMBERS

Oxidation Number Rules:

1. The oxidation number of any pure element is 0.
2. The oxidation number of a monatomic ion equals that charge on the ion.
3. The more electronegative element in a binary compound is assigned the number equal to the charge it would have if it were an ion.
4. The oxidation number of fluorine in a compound is always -1.
5. Oxygen has an oxidation number of -2 unless it is combined with F, in which it is +1 or +2, or it is in peroxide (such as H_2O_2 or Na_2O_2), in which it is -1.
6. Hydrogen is +1, unless combined with a metal, and then it is -1.
7. In compounds, Group 1 is +1, Group 2 is +2, and Aluminum is +3.
8. The sum of the oxidation numbers of all atoms in a neutral compound is 0.
9. The sum of the oxidation numbers in a polyatomic ion equals the charge of the ion.

Part A: In the following questions, give the oxidation number of the indicated atoms/ion:

- | | |
|---------------------------------------------|--------------------------------------------|
| 1. N in N_2O_5 _____ | 16. C in CH_4 _____ |
| 2. S in H_2SO_4 _____ | 17. Mn in MnO_2 _____ |
| 3. C _____ | 18. S in SO_3^{2-} _____ |
| 4. C in CO _____ | 19. Mg^{2+} _____ |
| 5. Na in NaCl _____ | 20. Cl _____ |
| 6. H in H_2O _____ | 21. O_2 _____ |
| 7. Ba in BaCl_2 _____ | 22. P _____ |
| 8. N in NO_2^- _____ | 23. Na in Na_2S _____ |
| 9. S in Al_2S_3 _____ | 24. S in H_2S _____ |
| 10. S in HSO_4^- _____ | 25. Ca^{2+} _____ |
| 11. Cl in $\text{Fe}(\text{ClO}_2)_3$ _____ | 26. C in CN^- _____ |
| 12. Fe in $\text{Fe}(\text{ClO}_2)_3$ _____ | 27. H in OH^- _____ |
| 13. N in NO_3^- _____ | 28. Mn in KMnO_4 _____ |
| 14. Cu^{2+} _____ | 29. I in $\text{Mg}(\text{IO}_3)_2$ _____ |
| 15. Zn^{2+} _____ | 30. C in $\text{C}_2\text{O}_4^{2-}$ _____ |

Oxidation Number Practice Answer Key

Disha Experts

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