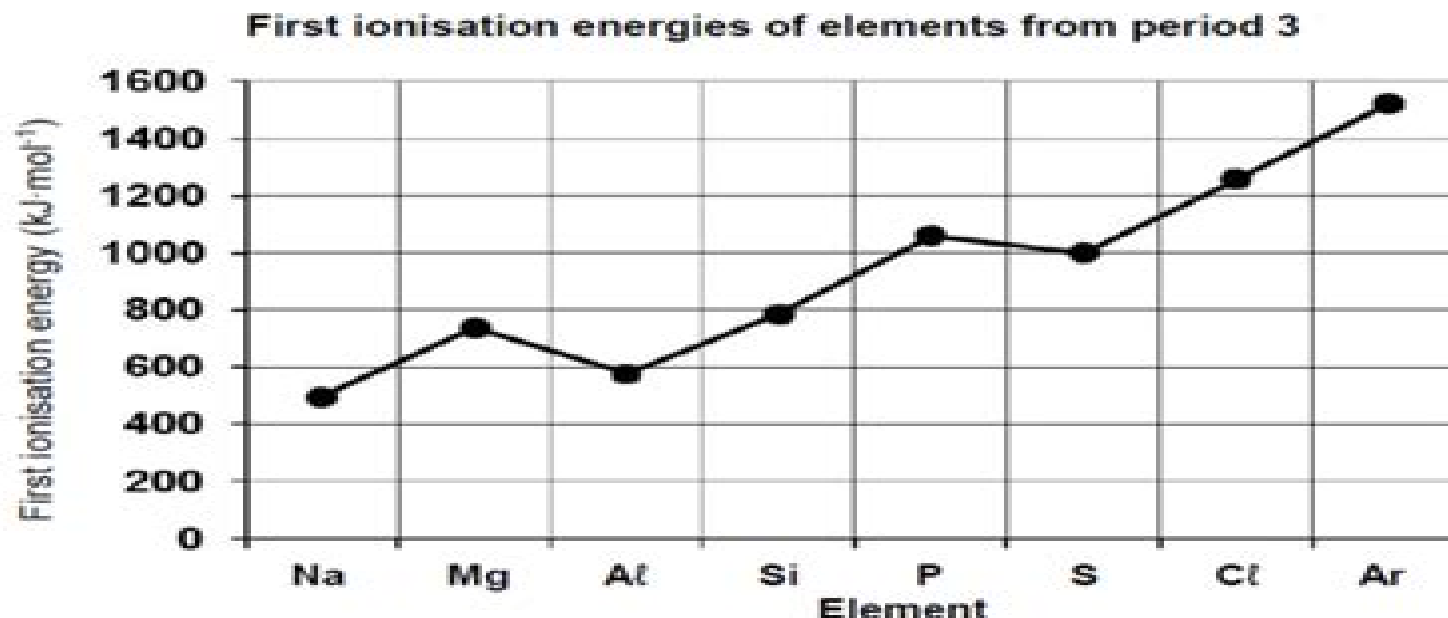


## QUESTION 6

The graph below shows the first ionisation energies of elements from period 3 in the periodic table.



- 6.1 Define the term *first ionisation energy*. (2)
- 6.2 Explain the difference in first ionisation energy between sodium and magnesium. (2)
- 6.3 From the graph, write down the approximate first ionisation energy of sodium. (1)
- 6.4 Explain why magnesium becomes ionised easily. (2)
- 6.5 Use the information in the graph above and calculate the energy needed to remove an electron from sodium atom. (3)
- 6.6 Draw an Aufbau diagram of a sodium ion. (2)
- 6.7 The Haber process uses the reaction between nitrogen and hydrogen to make ammonia. The formation of ammonia is exothermic, releasing 46 kilojoules per mole of ammonia formed.



- 6.7.1 What type of bond exists in the molecule of ammonia? (1)
- 6.7.2 Explain your answer to QUESTION 6.7.1. (2)

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