

ORGANIC CHEMISTRY

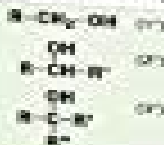
SECTION 1

CHEMISTRY OF CARBON

- Organic chemistry is based on carbon compounds, which form a wide variety of compounds, with H, O, N, S, and a few other elements.
- Ground state configuration of carbon is $1s^2 2s^2 2p^2$
- Easily hybridises to produce 4 sp^3 orbitals or 3 sp^2 orbitals and 1 p orbital or 2 sp orbitals and 2 p orbitals, carbon is tetravalent
- Carbon ring and chain structures may have single, double, and triple bonds between carbon atoms.
- Carbon can form 4 covalent bonds to other carbons, forming rings, branched or unbranched chains, isomerisation of rings and chains.
- Carbon compounds are relatively unstable to heat (i.e., decomposition at elevated temperatures)

ALCOHOLS

- Alcohols are organic compounds with the hydroxyl functional group ($-OH$) bonded to an hybridized carbon atom
- Classified as primary (1°), secondary (2°) or tertiary (3°) according to classification of carbon that bears the $-OH$ group



- Form hydrogen bonds, they have higher boiling points than esters or hydrocarbons of the same weight
- Short-chain alcohols are soluble in water
- Methanol, propanol, and butanol are very toxic
- Formulation: $C_nH_{2n+2}O$
 $\rightarrow C_nH_{2n+2}OH + 2CO_2$
(fermentation of sugar using yeast)



AMINES

- Amines are organic derivatives of ammonia (NH_3)
- Amino hydrogen atoms of ammonia are replaced by an alkyl or aryl groups, the result is primary (1°), secondary (2°) or tertiary (3°)
- Amines are pyramidal in shape
- May form hydrogen bonds, though weaker than bonds in alcohols
- Important in biological information, used in disinfectants, dyes, and insecticides
- Known as organic bases

- Amino acids are building blocks of proteins
- Synthesis: Reduction of nitrogen oxides and oxides, reduction, substitution, of nitriles and nitroalkenes and nitroalkenes
- Synthesis of 1° amines



ORGANIC REACTIONS

ACIDS, BASES, ELECTROPHILES & NUCLEOPHILES

- A Lewis acid is an atom, molecule or ion that can accept an electron pair
- A Brønsted-Lowry acid donates a proton
- All electrophiles, they are attracted to electron-rich areas
Examples: H^+ , $2NO_2$, NO_2^+ , NO_3^+ , NO_2^- , carbocations
- A Lewis base is an atom, molecule or ion that can donate an electron pair
- A Brønsted-Lowry base accepts a proton
- All nucleophiles, they are attracted to proton centers (nucleus)
Examples: OH^- , H_2O , CH_3O^- , ROH , CH_3O^- , CH_3^- , carbanions

ADDITION

- A reaction in which 2 molecules combine to yield one molecule, small ion fragments
- Addition frequently occurs at a double or triple bond and sometimes at nucleophile

Nucleophile:



Example: Nucleophile Y^- bonds to carbon of the $C=C$ group, covalent carbon (e.g., nucleophile) bonds to oxygen

Electrophile:



Markovnikov's Rule
 • When HX is added to an alkene, H attaches to carbon which has more H s attached
 • Does not always apply (e.g., addition of HBr with peroxide present)

ELIMINATION

- Reverse of addition, two atoms or groups are removed from a molecule



Example: A double or triple bond is formed by the loss of 2 atoms or groups of atoms

SUBSTITUTION

- A reaction in which an atom or group of atoms in a molecule or ion is replaced by another atom or group of atoms

Nucleophilic Substitution

- Nucleophile Y^- displaces group X (e.g., H , OH , Cl , Br)



Electrophilic Substitution

- An electrophile E^+ displaces a hydrogen atom in the aromatic ring



REARRANGEMENT

- Reaction that converts a species to a constitutional isomer by moving bonds around

Example: Intramolecular migration of a hydride or methyl group to an adjacent positively charged carbon results in a more stable carbocation



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Regenstein,2012-12-02 Food Protein Chemistry An Introduction for Food Scientists discusses food proteins and how they are studied Proteins are both biological entities and physicochemical compounds and they will be examined in both contexts in this volume The chemical and physical properties of proteins will be viewed from the perspective of chemists despite the fact that their use in the food supply emphasizes their biological nature Key topics discussed include proteins as essential to life amino acids protein classification selected proteins of the most important food systems and protein structure The book also includes chapters on protein measurement protein purification and spectral techniques for the study of proteins The book requires readers to have the equivalent of the Institute of Food Technologists requirements for undergraduate food science majors It also assumes a knowledge of math through calculus While primarily intended for senior and first year graduate food science students the text may also be useful to researchers in allied fields **The Grants Register 2024** Palgrave

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