Precipitation Reactions

For each of the following pairs of aqueous solutions, write a balanced molecular equation.

- Use your molecular equation to predict whether or not a precipitation reaction has occurred.
- If a reaction has occurred, write the ionic and net ionic reactions.
- Sodium hydroxide and calcium bromide
- Lead (II) nitrate and hydrochloric acid (HCI)
- 3. Sodium carbonate and potassium fluoride
- 4. Silver nitrate and copper (II) sulfate
- Silver nitrate and nickel (II) chloride

Answers:

- 2NaOH₁₀₀ + CaBr₂₀₀₁ + 2NaBr₂₀₀ + Ca(OH)₂₀₀
 Will this reaction occur? Yes
 2Na*₂₀₀₁ + 2OH*₂₀₀₁ + Ca*₂₀₀₁ + 2Br*₂₀₀₁ → 2Na*₂₀₀₁ + 2Br*₂₀₀₁ + Ca(OH)₂₀₀₁
 Ca*₂₀₀₁ + 2OH*₂₀₀₁ → Ca(OH)₂₀₀₁
- 2) Pb(NO₂)₂₅₀₀ + 2HCl₁₆₀ → 2HNO₂₅₋₆ + PbCl₂₆₀ Will this reaction occur? Yes Pb^{2*}₃₋₆ + 2NO₂₃₋₆ + 2H*₃₋₆₀ + 2Cl*₃₋₆ → 2H*₃₋₆₀ + 2NO₂₃₋₆ + PbCl₂₆₀ Pb^{2*}₃₋₆₀ + 2Cl*₃₋₆₀ → PbCl₂₆₀
- Na₂CO_{Set} + 2 KF_{bril} → 2 NaF_{bril} + K₂CO_{Set} Will a reaction occur? No
- 4) 2AgNO_{Stell} + CuSO_{Stell} + Ag₂SO_{Stell} + Cu(NO₃)_{Stell} *Ag₃SO_{Stell} is often listed as soluble Will this reaction occur? Yes

 2Ag*_{lock} + 2NO_{3 lock} + Cu^{2*}_{lock} + SO₃*_{lock} → Ag₃SO_{Stell} + Cu^{2*}_{lock} + 2NO_{3 lock}

 2Ag*_{lock} + SO₃*_{lock} → Ag₃SO_{Stell}
- 5) 2AgNO_{Set} + NiCl_{Set} + 2AgCl₁₆ + Ni(NO₃)_{Set}
 Will this reaction occur? Yes
 2Ag'_{Set} + 2NO_{3 Set} + Ni^{2*}_{Set} + 2Cl'_{Set} → 2AgCl₁₀ + Ni^{2*}_{Set} + 2NO_{3 Set}
 2Ag'_{Set} + 2Cl'_{Set} → 2AgCl₁₀ becomes Ag'_{Set} + Cl'_{Set} → AgCl₁₀ (need to reduce coeffecients)

Precipitation Reactions Answer Sheet

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