

Balancing Equations and Predicting Products

2 Part 1: Balancing Equations

Directions: Balance the following chemical equations and identify the type of reaction.

1. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ Type of reaction: Synthesis
2. $\text{Mg} + \text{HCl} \rightarrow \text{MgCl}_2$
Type of reaction: Acid base
3. $\text{Mg} + \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$
Type of reaction: Single displacement
4. $\text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}$
Type of reaction: Double displacement
5. $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
Type of reaction: Decomposition
6. $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
Type of reaction: Combustion

3 Part 2: Predicting Products

Directions: Given the reactants, predict the products of the following chemical reactions.

1. $\text{Na} + \text{Cl}_2 \rightarrow ?$
2. $\text{Ca}(\text{OH})_2 + \text{H}_2\text{SO}_4 \rightarrow ?$
3. $\text{Mg}(\text{NO}_3)_2 + \text{NaOH} \rightarrow ?$
4. $\text{HNO}_3 + \text{KOH} \rightarrow ?$

Predicting Products And Balancing Equations

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Fetzer-Gislason, 2003-03-21 The manual contains laboratory experiments written specifically for the prep chem lab as well as for the general chemistry course Available as a complete manual or custom published at <http://custompub.whfreeman.com>

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