

Revision - Principles of chemistry 1e - Chemical formulae and calculations

Where do you find the mass number on the periodic table? The big number

How do you calculate the mass of a NaOH: Add the top numbers together - $23+16+1 = 40$

a) F_2 _____

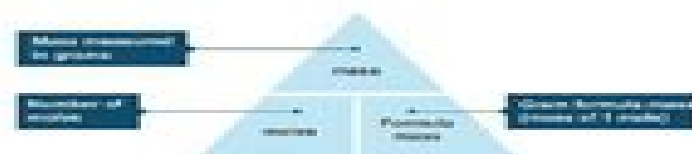
c) H_2SO_4 _____

e) $\text{Mg}(\text{OH})_2$ _____

b) Fe _____

d) Al_2O_3 _____

f) $\text{Al}(\text{NO}_3)_3$ _____



Calculate the number of moles in

a) 100 g of Cu _____

b) 1 kg of Al_2O_3 _____

c) 20 mg of Au _____

d) 0.35g HNO_3 _____

Calculate the mass of

a) 5 moles of N_2 _____

b) 0.02 moles of Al_2O_3 _____

c) 2 moles of NaCl _____

e) 0.05 moles of NaCl _____

d) 0.2 moles C_2H_6 _____

Percentage yield

A chemistry teacher burned 50g of copper in oxygen, hoping to make 80g of reaction only produced 67g of this product. What was the percentage yield?

copper oxide. However, the

$$\% \text{ Yield} = \frac{\text{Actual Yield}}{\text{Predicted Yield}} \times 100$$

Preliminary Chemistry Revision Questions

David Mahony



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