

## 11-2 Practice Problems

1. Determine the mass of a solution produced when 10.0 g of 1-butanol reacts with excess oxygen.

According to the following equation:  

$$C_4H_{10}O + 12O_2 \rightarrow 4CO_2 + 5H_2O$$

**Answer: 26.9 g  $CO_2$  + 13.1 g  $H_2O$**

2. Determine the mass of butane,  $C_4H_{10}$ , produced when 1.00 mol of butanol reacts with 10.0 g of hydrogen gas,  $H_2$ , according to the following equation:

**Answer: 1.00 mol  $C_4H_{10}$  + 1.00 mol  $H_2O$**

3. Determine the mass of butane,  $C_4H_{10}$ , produced when 10.0 g of hydrogen reacts with 1-butanol according to the following equation:

$2C_4H_9OH + 5H_2 \rightarrow 2C_4H_{10} + 4H_2O$

4. Determine the mass of butanol,  $C_4H_9OH$ , produced when 10.0 g of hydrogen,  $H_2$ , reacts with excess 1-butanol according to the following equation:

$2C_4H_9OH + 5H_2 \rightarrow 2C_4H_{10} + 4H_2O$

5. Determine the amount of hydrogen gas,  $H_2$ , required to produce 1.00 mol of butane,  $C_4H_{10}$ , according to the following equation:

$2C_4H_9OH + 5H_2 \rightarrow 2C_4H_{10} + 4H_2O$

6. Determine the amount of hydrogen gas,  $H_2$ , required to produce 10.0 g of butane,  $C_4H_{10}$ , according to the following equation:

$2C_4H_9OH + 5H_2 \rightarrow 2C_4H_{10} + 4H_2O$

7. Determine the mass of butanol,  $C_4H_9OH$ , produced when 10.0 g of hydrogen,  $H_2$ , reacts with excess 1-butanol according to the following equation:

$2C_4H_9OH + 5H_2 \rightarrow 2C_4H_{10} + 4H_2O$

8. Determine the mass of butanol,  $C_4H_9OH$ , produced when 1.00 mol of hydrogen,  $H_2$ , reacts with excess 1-butanol according to the following equation:

$2C_4H_9OH + 5H_2 \rightarrow 2C_4H_{10} + 4H_2O$

9. How many moles of hydrogen gas are necessary for the combustion of 10.0 g of butane, according to the reaction shown in the following equation?  
 $C_4H_{10} + 13O_2 \rightarrow 4CO_2 + 5H_2O$

10. Find the amount of oxygen,  $O_2$ , required to produce 1.00 mol of butane,  $C_4H_{10}$ , according to the following equation:  
 $2C_4H_9OH + 5H_2 \rightarrow 2C_4H_{10} + 4H_2O$

11. How many grams of butane react to produce 10.0 g of  $CO_2$  gas?

**Answer: 1.00 g  $C_4H_{10}$  + 1.00 g  $H_2O$**

12. Find the amount of butane,  $C_4H_{10}$ , required to produce 1.00 mol of  $CO_2$ , according to the following equation:  
 $2C_4H_9OH + 5H_2 \rightarrow 2C_4H_{10} + 4H_2O$

**Answer: 1.00 mol  $C_4H_{10}$  + 1.00 mol  $H_2O$**

# Prentice Hall Chemistry 11 2 Practice Problems Answers

**M Mosston**



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