

## 14

## THE BEHAVIOR OF GASES

## Chapter Test II

## A. Matching

Match a word or phrase in Column I with the correct description in Column II. Write the letter of the correct answer next to the item.

## Column I

1. The constant relating unit temperature changes to changes in volume for a fixed amount of gas at constant pressure (the ratio of the change in volume to the change in temperature).
2. The volume of a fixed amount of a gas at constant pressure in the standard state (0°C and 1 atm).
3. The ratio of the mass of a gas to its volume, independent of the actual amount of gas used.
4. An experimental gas constant determined from the ideal gas law at low pressures.
5. A constant of nature that has values of about 8.314 J/mol·K.
6. The a parameter of gas constant temperatures for which the gas molecules move with greatest freedom.
7. An experimental gas constant derived from the ideal gas law at low pressures.
8.  $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
9. A constant of nature through which Boyle's law can be expressed.
10. A constant of nature (8.314 J/mol·K).

## Column II

- a. standard pressure
- b. ideal gas constant (R)
- c. volume
- d. compressibility
- e. Boyle's law
- f. Boyle constant
- g. standard temperature
- h. density
- i. compressibility
- j. compressibility coefficient

## B. Multiple Choice

Choose the correct answer and write its letter on the line.

1. According to Graham's law, the rate of effusion of a gas is proportional to the square root of its molecular weight. Which of the following gases effuses most rapidly?  
a. the heaviest gas  
b. the lightest gas  
c. the gas with the highest molecular weight  
d. the gas with the lowest molecular weight

# Prentice Hall Chemistry Test Prep Answers

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Staley, Michael S. Matta, Edward L.  
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