

# 13.1 THE NATURE OF GASES

P.1

## Section Review

### Objectives

- Describe the assumptions of the kinetic theory as it applies to gases
- Interpret gas pressure in terms of kinetic theory
- Define the relationship between Kelvin temperature and average kinetic energy

### Vocabulary

- |                  |                        |                             |
|------------------|------------------------|-----------------------------|
| • kinetic energy | • vacuum               | • pascal (Pa)               |
| • kinetic theory | • atmospheric pressure | • standard atmosphere (atm) |
| • gas pressure   | • barometer            |                             |

### Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

The kinetic theory describes the 1 of particles in matter and the forces of attraction between them. The theory assumes that the volume occupied by a gas is mostly 2, that the particles of gas are relatively 3, move 4 of each other, and are in constant 5 motion. The 6 between particles are perfectly elastic so that the total 7 remains constant. Gas pressure results from the simultaneous collisions of billions of particles with an object. Barometers are used to measure 8 pressure. Standard conditions are defined as a temperature of 9 and a pressure of 10.

- motion
- empty space
- far apart
- independently
- random / rapid
- collisions
- Kinetic energy
- atmospheric
- 0°C / 273 K
- 1 atm / 101.3 kPa

### Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

ST 11. Atmospheric pressure is 760 mm Hg.

AT 12. The SI unit of pressure is the pascal.

↳ a multiple of kilo pascal

18 pts  
 - 2 pt if late  
 - 1 pt if NO ID

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Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science ,2003-11

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**Physical Chemistry for the Life Sciences** Thomas Engel,Gary Drobny,Philip J. Reid,2008 KEY BENEFIT Physical Chemistry for the Life Sciences presents the core concepts of physical chemistry with mathematical rigor and conceptual clarity and develops the modern biological applications alongside the physical principles The traditional presentations of physical chemistry are augmented with material that makes these chemical ideas biologically relevant applying physical principles to the understanding of the complex problems of 21st century biology KEY TOPICS Physical Chemistry Biology MARKET For all readers interested in physical chemistry and biology      **General Catalogue of Printed Books** British Museum. Department of Printed Books,1969      **Addison-Wesley Chemistry** ,2002

## **Prentice Hall Chemistry Worksheet Answers Chapter 13** Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Prentice Hall Chemistry Worksheet Answers Chapter 13**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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