

Nuclear Decay

Types of Radiation

1. Complete the table below to compare the properties of alpha, beta, and gamma radiation.

Name	Alpha	Beta	Gamma
Greek letter	α	β	γ
Symbol	${}^4_2\text{He}$	${}^0_{-1}e$	γ
Composition	2 protons, 2 neutrons	an electron	radiation (light)
Charge	0	-1	0
Stopped by	paper, skin, clothing	aluminum	lead

Nuclear Decay Reactions

In problems 2 – 5, fill in the missing particle for each nuclear decay reaction. Then, match each equation to the appropriate type of nuclear decay.

- _____ 2. ${}^{14}_6\text{C} \rightarrow {}^{14}_7\text{N} + \text{_____}$ beta A. alpha emission
- _____ 4. ${}^{218}_{84}\text{Po} \rightarrow {}^{214}_{82}\text{Pb} + \text{_____}$ alpha B. beta emission
- _____ 5. $\text{_____} + {}^0_{-1}e \rightarrow {}^{147}_{60}\text{Nd}$ electron capture C. electron capture

In problems 6 and 7, write a balanced equation for each nuclear decay reaction.

6. Decay of radium-223 by alpha (α) emission.



7. Decay of iodine-126 by electron (β) emission.



8. How are the atomic number and mass number of a nuclide affected by Beta Decay?

The mass number does not change but the atomic number increases by 1.

9. How are the atomic number and mass number of a nuclide affected by the emission of gamma radiation?

The atomic number and mass number do not change.

10. Which has a slightly greater mass in a nuclear reaction, the reactants or the products? Why?

It depends on the type of reaction. A fusion reaction will have greater mass in the products because it is combining two nuclei. A fission reaction will not change mass.

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