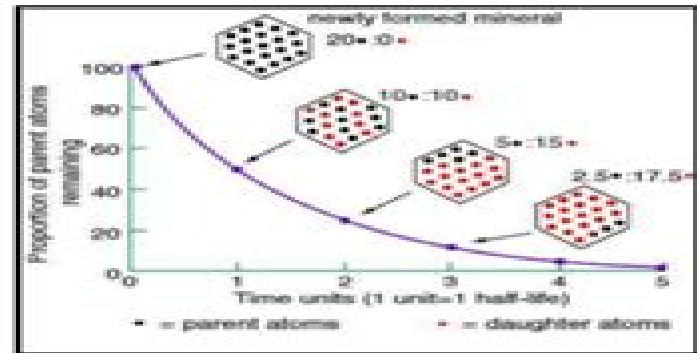


Absolute Age and the Radioactive Dating of Rocks

Rocks are a form of matter. All the matter you see, including rocks, is made of tiny particles called atoms. When all the atoms of a particular type of matter are the same, the matter is an element.

Most elements are **stable**. They do not change under normal conditions. But some elements exist in forms that are **unstable**. Over time, these elements break down, or decay, by releasing particles and energy in a process called **radioactive decay**. These unstable elements are said to be **radioactive**.

During radioactive decay, the atoms of one element break down to form atoms of another element. Radioactive elements occur naturally in igneous rocks (rocks formed from magma). For an igneous rock, its "birthday" is when it first hardens to become rock. As a radioactive element within the igneous rock decays, it changes into another element. Therefore, the composition of the rock changes slowly over time. The amount of the radioactive element (parent) goes down. But the amount of the new element (daughter) goes up. The rate of decay of each radioactive element is constant; it never changes. This rate of decay is the element's half-life. The half-life of a radioactive element is the time it takes for half of the radioactive atoms (parent) to decay, or change into a stable element (daughter). Scientists use this rate of decay to calculate a rock's age.



Geologists use radioactive dating to determine the absolute age of rocks. *The absolute age of a rock is its age in numbers of years.* In radioactive dating, scientists first determine the amount of a radioactive element in a rock. Then they compare that amount with

the amount of the stable element into which the radioactive element decays. So if a rock has 50% parent element X and 50% daughter element Z, and the half life of the X is 100 years, the rock would be 100 years old. If it was 25 % X and 75% Z that means it went through 2 half lives and is 200 years old.

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