

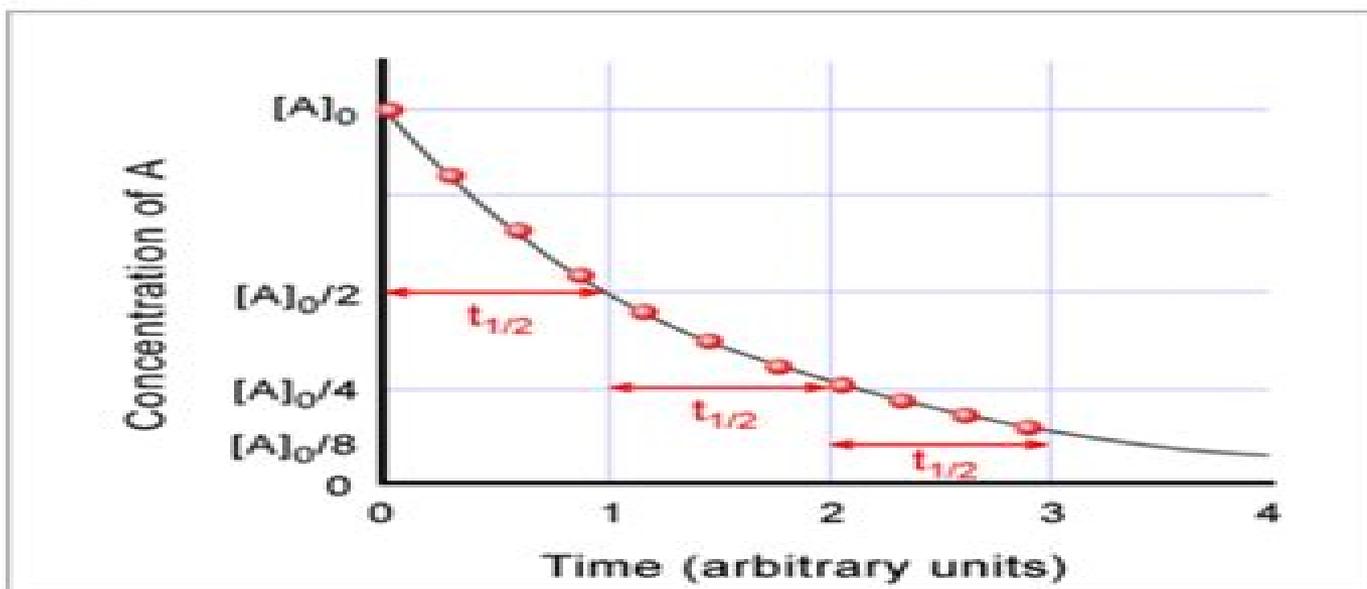
Penny Decay

Radioactive decay follows 1st order kinetics and in the reaction, the concentration of the reactant decreases exponentially. The rate of the reaction equals the concentration of the reactant, $[A]$, raised to the first power times a proportionality constant, k , which is called the rate constant. The rate constant is a fixed value for a given reaction.

$$\text{Rate} = k [A]$$

$[A]$ is the concentration of the reactant and k is the rate constant.

A plot of the concentration of A versus time is show below. You can see that the concentration of A does not decrease in a linear manner. The concentration rapidly decreases at the beginning but slows down over time. This is characteristic of an exponential decrease.



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K Payea



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