

Name Key Class \_\_\_\_\_ Date \_\_\_\_\_

## Nuclear Decay

### Types of Radiation

1. Complete the table below to compare the properties of alpha, beta, positron, and gamma radiation.

Name	Alpha	Beta	Positron	Gamma
Greek letter	$\alpha$	$\beta^-$	$\beta^+$	$\gamma$
Symbol	${}_2^4\text{He}$	${}_{-1}^0e$	${}_{+1}^0e$	N/A
Composition	$2p^+$ , $2n^0$	electron	anti-electron	EM wave
Charge	$2+$	$1-$	$1+$	$0$
Stopped by	paper, skin, clothing	wood, glass	wood, glass	thick concrete or lead

### Nuclear Decay Reactions

In problems 2 – 5, fill in the missing particle for each nuclear decay reaction. Then, match each equation to the appropriate type of nuclear decay.



A. alpha emission



B. beta emission



C. electron capture



D. positron emission

In problems 6 and 7, write a balanced equation for each nuclear decay reaction.

6. Decay of radium-223 by alpha ( $\alpha$ ) emission.



7. Decay of iodine-126 by positron ( $\beta^+$ ) emission.



8. How are the atomic number and mass number of a nuclide affected by positron emission?

Atomic number decreases by one. Mass number stays the same.

9. How are the atomic number and mass number of a nuclide affected by the emission of gamma radiation?

Both atomic number and mass number stay the same since gamma radiation has no mass or charge.

10. Which has a slightly greater mass in a nuclear reaction, the reactants or the products? Why?

The reactants have a slightly greater mass. In a nuclear reaction, a small amount of mass is converted to energy according to the equation  $E = mc^2$ . The difference in mass is referred to as the mass defect.

# Radioactivity And Nuclear Reaction Worksheet Answers

**Niti Arora**

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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## **Table of Contents Radioactivity And Nuclear Reaction Worksheet Answers**

1. Understanding the eBook Radioactivity And Nuclear Reaction Worksheet Answers
  - The Rise of Digital Reading Radioactivity And Nuclear Reaction Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Radioactivity And Nuclear Reaction Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Radioactivity And Nuclear Reaction Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Radioactivity And Nuclear Reaction Worksheet Answers
  - Personalized Recommendations
  - Radioactivity And Nuclear Reaction Worksheet Answers User Reviews and Ratings
  - Radioactivity And Nuclear Reaction Worksheet Answers and Bestseller Lists
5. Accessing Radioactivity And Nuclear Reaction Worksheet Answers Free and Paid eBooks
  - Radioactivity And Nuclear Reaction Worksheet Answers Public Domain eBooks
  - Radioactivity And Nuclear Reaction Worksheet Answers eBook Subscription Services
  - Radioactivity And Nuclear Reaction Worksheet Answers Budget-Friendly Options

6. Navigating Radioactivity And Nuclear Reaction Worksheet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Radioactivity And Nuclear Reaction Worksheet Answers Compatibility with Devices
  - Radioactivity And Nuclear Reaction Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Radioactivity And Nuclear Reaction Worksheet Answers
  - Highlighting and Note-Taking Radioactivity And Nuclear Reaction Worksheet Answers
  - Interactive Elements Radioactivity And Nuclear Reaction Worksheet Answers
8. Staying Engaged with Radioactivity And Nuclear Reaction Worksheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Radioactivity And Nuclear Reaction Worksheet Answers
9. Balancing eBooks and Physical Books Radioactivity And Nuclear Reaction Worksheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Radioactivity And Nuclear Reaction Worksheet Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Radioactivity And Nuclear Reaction Worksheet Answers
  - Setting Reading Goals Radioactivity And Nuclear Reaction Worksheet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Radioactivity And Nuclear Reaction Worksheet Answers
  - Fact-Checking eBook Content of Radioactivity And Nuclear Reaction Worksheet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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