

## SEMESTER 2 STUDY GUIDE

**You must show your work for any question that requires you to do calculations**

### Unit 8 and Unit 9 - The Mole and Stoichiometry

1. Give a definition of the term mole  
a specific number of particles, which is approximately  $6.022 \times 10^{23}$  particles, also known as Avogadro's number.
2. How many liters of a gas are in one mole?  
one mole of any ideal gas occupies 22.4 liters.
3. How many particles/atoms/molecules/formula units are in one mole?  
 $6.022 \times 10^{23}$  particles.
4. How do you determine one mole when it comes to the mass of a substance?  
(g/mol)
5. Convert 2 moles of  $O_2$  gas into mass in grams.  
32 g/mol. Therefore, 2 moles of  $O_2$  gas is equal to 64 grams of  $O_2$  gas.
6. Convert 3.5 moles of NaCl into formula units.  
 $6.022 \times 10^{23}$  particles/mol. Therefore, 3.5 moles of NaCl is equal to  $2.11 \times 10^{24}$  formula units of NaCl.
7. Convert 0.25 moles of chlorine gas into liters.  
 $V = nRT/P = (0.25 \text{ mol})(0.0821 \text{ L}\cdot\text{atm/mol}\cdot\text{K})(273 \text{ K})/(1 \text{ atm}) = 5.57$  liters of chlorine gas.
8. Convert 56.25 grams of water into moles.  
18 g/mol. Therefore, 56.25 grams of water is equal to 3.125 moles of water.
9. Convert 15.65 liters of  $N_2$  gas into moles.  
 $n = PV/RT = (1 \text{ atm})(15.65 \text{ L})/(0.0821 \text{ L}\cdot\text{atm/mol}\cdot\text{K})(273 \text{ K}) = 0.680$  moles of  $N_2$  gas.
10. Convert  $1.65 \times 10^{24}$  atoms of Mg into moles.  
 $6.022 \times 10^{23}$  particles/mol. Therefore,  $1.65 \times 10^{24}$  atoms of Mg is equal to 2.74 moles of Mg.
11. Convert  $1.65 \times 10^{24}$  atoms of Mg to mass.  
24.31 g/mol. Therefore,  $1.65 \times 10^{24}$  atoms of Mg is equal to 667 grams of Mg.
12. Convert 6.25 grams of  $CO_2$  to molecules of  $CO_2$   
 $(0.142 \text{ mol})(6.022 \times 10^{23} \text{ molecules/mol}) = 8.56 \times 10^{22}$  molecules of  $CO_2$ .
13. Determine the limiting reactant if you have a 3.86 grams sample of  $NO_2$  and a 45.25 gram sample of  $O_2$ .  
 $2NO_2(g) + O_2 \rightarrow 2NO$   
 $(3.86 \text{ g})/(46.01 \text{ g/mol}) = 0.084 \text{ mol of } NO_2$
14. What is the maximum amount of NO that can be produced from the reaction in #13?  
 $2NO_2(g) + O_2(g) \rightarrow 2NO(g)$

# Slo Study Guide Semester 2

**Judd E. Hollander**



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## **Table of Contents Slo Study Guide Semester 2**

1. Understanding the eBook Slo Study Guide Semester 2
  - The Rise of Digital Reading Slo Study Guide Semester 2
  - Advantages of eBooks Over Traditional Books
2. Identifying Slo Study Guide Semester 2
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Slo Study Guide Semester 2
  - User-Friendly Interface
4. Exploring eBook Recommendations from Slo Study Guide Semester 2
  - Personalized Recommendations
  - Slo Study Guide Semester 2 User Reviews and Ratings

- Slo Study Guide Semester 2 and Bestseller Lists
- 5. Accessing Slo Study Guide Semester 2 Free and Paid eBooks
  - Slo Study Guide Semester 2 Public Domain eBooks
  - Slo Study Guide Semester 2 eBook Subscription Services
  - Slo Study Guide Semester 2 Budget-Friendly Options
- 6. Navigating Slo Study Guide Semester 2 eBook Formats
  - ePub, PDF, MOBI, and More
  - Slo Study Guide Semester 2 Compatibility with Devices
  - Slo Study Guide Semester 2 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Slo Study Guide Semester 2
  - Highlighting and Note-Taking Slo Study Guide Semester 2
  - Interactive Elements Slo Study Guide Semester 2
- 8. Staying Engaged with Slo Study Guide Semester 2
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Slo Study Guide Semester 2
- 9. Balancing eBooks and Physical Books Slo Study Guide Semester 2
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Slo Study Guide Semester 2
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Slo Study Guide Semester 2
  - Setting Reading Goals Slo Study Guide Semester 2
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Slo Study Guide Semester 2
  - Fact-Checking eBook Content of Slo Study Guide Semester 2
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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