

crust
la croûte
die Kruste
la corteza
la crosta

white bread • le pain blanc
• das Weißbrot • el pan
blanco • il pane bianco



loaf
le pain
der Laib
la hogaza
il filone

brown bread • le pain bis
• das Graubrot • el pan
moreno • il pane nero



slice
la tranche
die Scheibe
la rebanada
la fetta

wholemeal bread • le pain de
son • das Vollkornbrot • el pan
integral • il pane integrale



granary bread • le pain
complet • das Mehrkornbrot
• el pan multicereales
• il pane di granoio



corn bread • le pain de maïs
• das Maisbrot • el pan de
maíz • il pane di maïs



soda bread • le pain à la
bicarbonate de soude
• das Sodabrot • el pan al
bicarbonato sódico • il pane
lievitato con bicarbonato di sodio



sourdough bread • le pain au
levain • das Sauerteigbrot
• el pan de masa madre
• il pane aspro



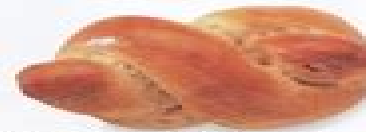
flatbread • le pain plat
• das Fladenbrot • el pan sin
levadura • la schiacciata



bagel • le petit pain américain
• das Hefebrotchen • la rosca
• la ciambella



bap • le petit pain rond • das
weiche Brötchen • el bollo • il
panino



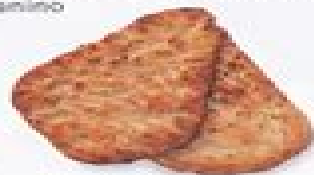
roll • le petit pain • das
Brötchen • el panecillo
• il filoncino



fruit bread • le pain aux
raisins secs • das Früchtebrot
• el plumcake • il pane alla
frutta



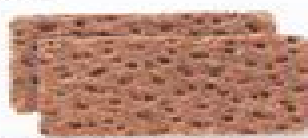
seeded bread • le pain aux
graines • das Körnerbrot • el pan
con semillas • il pane con semi



naan bread • le naan
• der Naan • el naan
• il naan



pitta bread • le pita
• das Pitabrot • el pan de
pita • la pita



crispbread • le biscuit
scandinave • das Knäckebrot
• el pan crujiente • i crackers

self-raising flour
la farine avec la levure
das Mehl mit Backpulver
la harina con levadura
la farina autolievitante

plain flour
la farine sans levure
das Mehl ohne Backpulver
la harina sin levadura
la farina semplice

prove (v)
lever
gehen lassen
levar
lievitare

breadcrumbs
la chapelure
das Paniermehl
el pan rallado
le briciole

slicer
la machine à couper
der Brotschneider
el rebanador
l'affettatrice

strong flour
la farine traitée
das angereicherte Mehl
la harina para pan
la farina per il pane

rise (v)
se lever
aufgehen
subir
crescere

glaze (v)
glacer
glasieren
glasear
glassare

flute
la flûte
das Stangenweißbrot
la barra
il filoncino

baker
le boulanger
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web the ions which do not participate in chemical reactions and present the same on both sides of the reactions are called spectator ions in the net chemical reaction the spectator ions are cancelled from both sides of the equation in other words spectate means watch the other ions react in an aqueous solution

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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2\text{Ag}^+ \text{aq} + \text{CrO}_4^{2-} \text{aq} \rightarrow \text{Ag}_2\text{CrO}_4 \text{s}$

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web chemical equation in which all dissolved ionic reactants and products including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are

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web jan 29 2020 the ionic form of this reaction is $2\text{Na}^+ \text{aq} + 2\text{Cl}^- \text{aq} + \text{Cu}^{2+} \text{aq} + \text{SO}_4^{2-} \text{aq} \rightarrow 2\text{Na}^+ \text{aq} + \text{SO}_4^{2-} \text{aq} + \text{CuCl}_2 \text{s}$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web mar 11 2014 solution ionic equation $\text{H}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{Na}^+ \text{aq} + \text{HCO}_3^- \text{aq} \rightarrow \text{Na}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$ the spectator ions are Cl^- and Na^+ net ionic equation $\text{H}^+ \text{aq} + \text{HCO}_3^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$ here is a video on spectator ions *chapter 8 04 ionic equations chemistry libretexts* - Mar 06 2023

web the six $\text{NO}_3^- \text{aq}$ ions and the six $\text{Na}^+ \text{aq}$ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{Ba}^{2+} \text{aq} + 2\text{PO}_4^{3-} \text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2 \text{s}$

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+ \text{OH}^- \rightarrow \text{H}_2\text{O}$ the spectator ions in KOH and HNO_3 are K^+ and NO_3^- how to identify and list the

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web the Ag^+ and Cl^- ions become AgCl but the Na^+ ions and the NO_3^- ions stay as Na^+ ions and NO_3^- ions these two ions are examples of spectator ions that do nothing in the overall course of a chemical reaction

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