



SAT PRACTICE ANSWER SHEET

COMPLETE MARK ●

EXAMPLES OF
INCOMPLETE MARKS



It is recommended that you use a No. 2 pencil. It is very important that you fill in the entire circle darkly and completely. If you change your response, erase as completely as possible. Incomplete marks or erasures may affect your score.

TEST NUMBER

SECTION 1

ENTER TEST
NUMBER

For instance, for Practice
Test #1, fill in the circle
for 0 in the first column
and for 1 in the second
column.

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BM King



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web a which chemical is the limiting reactant b how many grams of zns will be formed c how many grams of the excess reactant will remain after the reaction is over 3 which element is in excess when 3 00 grams of mg is ignited in 2 20 grams of pure oxygen what mass is in excess what mass of mgo is formed 4

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web reactant present in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated molecular equation chemical equation in which all reactants and products are represented as neutral substances

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web the key to recognizing which reactant is the limiting reagent is based on a mole mass or mass mass calculation

whichever reactant gives the lesser amount of product is the limiting reagent what we need to do is determine an amount of one product either moles or mass assuming all of each reactant reacts

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web the reactant which is in a lesser amount than is required by stoichiometry is the limiting reactant in an alternate method of finding the limiting reagent the amount of product formed by each reactant is calculated the limiting reactant is the reactant from which the minimum amount of product is formed also if we calculate the amount of

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limits how much product can be formed as we saw in example 1 there are many different ways to determine the limiting reactant but they all involve using mole ratios from the balanced chemical equation

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