

NCERT Solutions for Class 11 Chemistry

Chapter - 1 Some basic concepts of Chemistry

- Calculate the molecular mass of the following:
 - i. H₂O

Ans: The molecular mass of water, H₂O, can be calculated by the following steps given below:

= (2 x Atomic mass of hydrogen element) + (1 x Atomic mass of oxygen element)

- $= [(2 \times 1.0084 \text{ u}) + (1 \times 16.00 \text{ u})]$
- = 2.016 u + 16.00 u
- = 18.016 u or 18.02 u
- ii. CO.

Ans: The molecular mass of carbon dioxide, CO, is calculated below:

- = (1 × Atomic mass of carbon) + (2 × Atomic mass of oxygen)
- $= [(1 \times 12.011 \text{ u}) + (2 \times 16.00 \text{ u})]$
- = 12.011 u + 32.00 u
- = 44.01 u
- iii. CH,

Ans: The molecular mass of methane, CH₄, is calculated below in step by step manner:

- = (1 × Atomic mass of carbon) + (4 × Atomic mass of hydrogen)
- $= [(1 \times 12.011 \text{ u}) + (4 \times 1.008 \text{ u})]$
- = 12.011 u +4.032 u
- = 16.043 u
- Calculate the mass percent of different elements present in sodium sulphate(Na₂SO₄).

Ans: The given compound in the question is sodium sulphate and its formula is Na₂SO₄. Its molecular formula is calculated below:

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