

Science

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(Chapter – 1) (Chemical Reactions and Equations) (Exercises) (Class 10)

Question 1:

Which of the statements about the reaction below are incorrect?



- (a) Lead is getting reduced.
- (b) Carbon dioxide is getting oxidised.
- (c) Carbon is getting oxidised.
- (d) Lead oxide is getting reduced.

(i) (a) and (b)

(ii) (a) and (c)

(iii) (a), (b) and (c)

(iv) all

Answer 1:

(i) (a) and (b)

Question 2:

The above reaction is an example of a



- (a) combination reaction.
- (c) decomposition reaction.

- (b) double displacement reaction.
- (d) displacement reaction.

Answer 2:

(d) displacement reaction.

Question 3:

What happens when dilute hydrochloric acid is added to iron filings? Tick the correct answer.

- (a) Hydrogen gas and iron chloride are produced.
- (c) No reaction takes place.

- (b) Chlorine gas and iron hydroxide are produced.
- (d) Iron salt and water are produced.

Answer 3:

(a) Hydrogen gas and iron chloride are produced.

Question 4:

What is a balanced chemical equation? Why should chemical equations be balanced?

Answer 4:

A reaction which has an equal number of atoms of all the elements on both sides of the chemical equation is called a balanced chemical equation.

The law of conservation of mass states that mass can neither be created nor destroyed. Hence, in a chemical reaction, the total mass of reactants should be equal to the total mass of the products. It means that the total number of atoms of each element should be equal on both sides of a chemical equation. Hence, it is for this reason that chemical equations should be balanced.

Question 5:

Translate the following statements into chemical equations and then balance them.

- (a) Hydrogen gas combines with nitrogen to form ammonia.
- (b) Hydrogen sulphide gas burns in air to give water and sulphur dioxide.
- (c) Barium chloride reacts with aluminium sulphate to give aluminium chloride and a precipitate of barium sulphate.
- (d) Potassium metal reacts with water to give potassium hydroxide and hydrogen gas.

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