

CHEMISTRY Terminology

The ABCs of chemistry

A

A: Symbol for atomic mass number
absolute zero: Lowest temperature; 0 Kelvin
Ae: Symbol for the element actinium, $Z = 89$
accuracy: The closeness of measurement to true value
acetate anion: $C_2H_3O_2^-$, charge -1
acetic acid: $C_2H_4O_2$
acetone: Dimethyl ketone; organic solvent
acid-base indicator: A chemical that changes color if the pH changes
acid-base titration: A method for determining acid or base concentrations
acid definition: See Arrhenius acid, Brønsted-Lowry acid, Lewis acid
acid ionization constant (K_a): Equilibrium constant for acid dissociation
acidic solution: pH below 7
actinide: Element with $Z = 90$ or above; radioactive
activation energy (E_a): A process's energy barrier



adhesion: Attraction of unlike molecules or materials
Ag: Symbol for the element silver, $Z = 47$
Al: Symbol for the element aluminum, $Z = 13$
alcohol: Organic compound with -OH group; ROH
aldehyde: Carbonyl group bonded to hydrogen and an organic group
alkali metal family: Lithium, sodium, potassium, rubidium, cesium, and francium; column #1
alkaline earth metal family: Beryllium, magnesium, calcium, strontium, barium, and radium; column #2
alkane: Hydrocarbon; all C-C single bonds
alkene: Hydrocarbon; 1 or more C=C double bonds
alkyne: Hydrocarbon; 1 or more C≡C triple bonds
allotropes: Two or more forms of an element
alloy: Solution of 2 or more metals
alpha (α): Greek letter that denotes radioactive particle and various scientific variables
alpha particle: Helium nucleus; charge +2
Au: Symbol for the element auriferous, $Z = 79$
amalgam: Alloy of mercury with other metal
amine: Organic base; RRRN; modified ammonia
amino acid: Compound with organic acid and organic base property; forms proteins and peptides
ammonia: NH_3 , base
ammonium cation: NH_4^+ , charge +1
ammonium hydroxide: NH_4OH
amphoteric oxide: Exhibits both acid and basic properties
amu: Atomic mass unit; $1/12$ mass of C-12
anion: Ion with a negative charge
anode: Electrode that supports oxidation
antibonding MO: MO is less stable than separate AOs
AO: Abbreviation for atomic orbital

aq: Aqueous; "dissolved in water"
Ar: Symbol for the element argon, $Z = 18$
aromatic: Organic compound with a benzene ring
Arrhenius acid: Produces hydronium ion in water solution
Arrhenius base: Yields hydroxide ion in water solution
Arrhenius equation for rate constant (k):

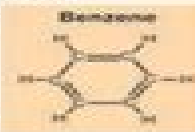
$$k = Ae^{-E_a/RT}$$
 E_a = activation energy
arsenide anion: As^{3-} , charge -3
As: Symbol for the element arsenic, $Z = 33$
At: Symbol for the element astatine, $Z = 85$
atm: Symbol for pressure in "atmospheres"
atom: Fundamental unit of all matter
atomic mass number (A): Total number of protons and neutrons
atomic number (Z): Number of protons in the nucleus
atomic orbital: Wave motion of electrons in atoms
atomic radius: Empirical measure of atom size
atomic weight: Weighted average of natural isotopes of an element
Au: Symbol for the element gold, $Z = 79$
Aufbau principle: Guides the filling of electronic subshells of the elements
Avogadro's law: Volume is proportional to gas moles at fixed pressure and temperature
Avogadro's number: $N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$

Bk: Symbol for the element berkelium, $Z = 97$
bleach: Chemical that is a strong oxidizing agent
boiling: Liquid \rightarrow gas at the boiling point
boiling point elevation: A solution has a higher boiling point than a pure solvent
boiling point (T_b): Liquid-gas equilibrium, $P = 1 \text{ atm}$
boiling point of water: 100°C
bond energy: Energy held by a chemical bond
bond length: Distance between 2 atoms in a bond
bond order for Lewis model: # of bonds divided by # of bonded atoms
bond order for MO treatment: # of filled bonding MOs minus # of filled antibonding MOs
bonding electrons: Form chemical bonds in a compound
bonding MO: MO is more stable than AOs
Boyle's law: PV constant for gas, fixed T and n
Br: Symbol for the element bromine, $Z = 35$
bromide anion: Br^- , charge -1
Brønsted-Lowry acid: Proton donor
Brønsted-Lowry base: Proton acceptor
buffer: Composed of weak acid and weak base; serves to keep pH constant



B

B: Symbol for the element boron, $Z = 5$
Ba: Symbol for the element barium, $Z = 56$
balanced equation: Each side of the equation has the same number of atoms of each element and charge
balanced redox: Electron loss = electron gain
barium cation: Ba^{2+} , charge +2
base definition: See Arrhenius base, Brønsted-Lowry base, Lewis base
base ionization constant (K_b): Equilibrium constant for base dissociation
basic solution: pH over 7
battery: Source of electrical power; galvanic cell
Be: Symbol for the element beryllium, $Z = 4$
benzene: C_6H_6 ; organic solvent; aromatic ring of 6 carbon atoms
beryllium cation: Be^{2+} , charge +2
beta (β): Greek letter that denotes radioactive particle and a number of scientific variables
beta particle: Energetic electron from the nucleus; charge -1
Bh: Symbol for the element bohrium, $Z = 107$
Bi: Symbol for the element bismuth, $Z = 83$
bicarbonate anion: HCO_3^- , charge -1
bimolecular oxygen: O_2
binary compound: Formed from 2 elements
bisulfate anion: HSO_4^- , charge -1



b: Symbol for centi- , the SI prefix for $1/100$ or 10^{-2}
c: Symbol for the speed of light
°C: Celsius temperature scale
C: Symbol for the element carbon, $Z = 6$
 $C_2H_4O_2$: Acetic acid
Ca: Symbol for the element calcium, $Z = 20$
calcium cation: Ca^{2+} , charge +2
calcium fluoride: CaF_2
calcium hydroxide: $Ca(OH)_2$
calorie: English unit of heat energy
calorimetry: Study of the release or absorption of heat
carbide anion: C^{4-} , charge -4
carbohydrate: Organic compound; carbon bonded to several -OH groups (sugar and starch)
carbon dioxide: CO_2
carbon monoxide: CO
carbonate anion: CO_3^{2-} , charge -2
carbonic acid: H_2CO_3
carbonyl group: CO , ligand complex
carboxyl group: $>CO$ bonded to 2 groups
carboxylic acid: Organic acid; $R-COOH$
catalyst: Accelerates a reaction but is not consumed in the reaction; lowers the activation energy
cathode: Electrode that supports reduction
cation: Ion with a positive charge
Cd: Symbol for the element cadmium, $Z = 48$



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Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises *Cracking the Virginia SOL* Tina Walton, Princeton Review (Firm), 2000 The Princeton Review realizes that acing the English Reading Literature and Research exam is very different from getting straight As in school The Princeton Review doesn't try to teach students everything there is to know about English only the techniques they'll need to score higher on the exam There's a big difference In *Cracking the Virginia SOL EOC English Reading Literature and Research* The Princeton Review will teach test takers how to think like the test makers and Eliminate incorrect answer choices by using Process of Elimination and other techniques Familiarize students with every type of reading material that might be on the test Use the two pass system to get a grip on the test structure and raise scores Hone individual skills with practice questions for each type of test question This book includes 2 full length simulated exams All of TPR's sample test questions are just like the ones test takers will see on the actual End of Course English Reading Literature and Research exam and TPR fully explains every solution Contents Include The Mystery Test Structure and Strategies II Terms Literary Movements and Writers Literature Review III Literary Elements and Printed and Resource Materials The Elements of Literature A Variety of Printed Material A Variety of Resource Material IV The Princeton Review Practice Tests **A Self-study Guide to the Principles of Organic Chemistry** Jiben Roy, 2013 A Self Study Guide to the Principles of Organic Chemistry Key Concepts Reaction Mechanisms and Practice Questions for the Beginner will help students new to organic chemistry grasp the key concepts of the subject quickly and easily as well as build a strong foundation for future study Starting with the definition of atom the author explains molecules electronic configuration

bonding hydrocarbons polar reaction mechanisms stereochemistry reaction varieties organic spectroscopy aromaticity and aromatic reactions biomolecules organic polymers and a synthetic approach to organic compounds The over one hundred diagrams and charts contained in this volume will help students visualize the structures and bonds as they read the text and make the logic of organic chemistry clear and easily understood Each chapter ends with a list of frequently asked questions and answers followed by additional practice problems Answers are included in the Appendix Self Study Guide CLAT 2022 Arihant Experts,2021-10-14 1 The Self Study Guide for the preparation of CLAT 2021 2 The book is divided into 6 Sections 3 Good number of MCQs have been provided for practice 4 Chapterwise Previous Years Papers and 3 Crack sets are given for thorough practice 5 Authentic Solutions are given for the complete assistance The Self Study Guide CLAT 2021 is a comprehensive textbook designed to give complete assistance for the preparation This book divides the entire syllabus into sections with Chapterwise theories along with sufficient number of MCQs are given for quick revision of topics making it a complete success package Each chapter carries previous years question from 2017 to 2011 for practice and 3 Practice sets are provided at end of the book to analyze the level of preparation With authentic solutions provided help students to strengthen the concepts Students who desire to score well in CLAT this self study guide is a perfect choice for them Toc CLAT Solved Paper 2020 2017 Legal Aptitude Indian Constitution English Language Mathematics Logical Reasoning General Knowledge Practice Sets 1 3 A Study Guide for Jose Raul Bernardo's "Silent Wing" Gale, Cengage Learning,2016 A Study Guide for Jose Raul Bernardo s Silent Wing excerpted from Gale s acclaimed Literature of Developing Nations for Students This concise study guide includes plot summary character analysis author biography study questions historical context suggestions for further reading and much more For any literature project trust Literature of Developing Nations For Students for all of your research needs *Study Guide to Accompany Calculus for the Management, Life, and Social Sciences* Clyde Metz,1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences Self Study Guide CLAT 2020 Arihant Experts,2019-08-12 CLAT Common Law Admission Test is an All India entrance examination and conducted by the National Law Schools Universities for the admission in the Under graduate and Post graduate programmes This examination is given right after 12 th grade for Undergraduate programmes in Law and after graduation in Law for Masters of Law This Self Study Guide CLAT 2020 book comprehensively covers the whole syllabus regarding to test and helps students to get through exam without any difficulty This book is divided into 6 parts namely Legal Aptitude Indian Constitution English Language General Knowledge Logical Reasoning and Mathematics This book provides Chapter wise theories with adequate amount of MCQs for recalling the concepts of each chapter Previous Years question papers have been provided in the chapter practice 3 sets of practice papers are provided at the end of the book which describes the exact pattern of the question to the candidates so that they can answer easily detailed explanation of solved papers 2019 2018 also have been provided for practice All the practice sets are also available for free on the WEB as well as on MOBILE so that

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and their applications It shows the reader how different types of materials are produced why they possess specific properties and how they are used in technology Each chapter contains study questions to enable discussions and consolidation of the acquired knowledge The new edition of this textbook is completely revised and updated to reflect the significant expansion of the field of materials chemistry over the last years covering now also topics such as graphene nanotubes light emitting diodes extreme photolithography biomedical materials and metal organic frameworks From the reviews of the first edition This book is not only informative and comprehensive for a novice reader but also a valuable resource for a scientist and or an industrialist for new and novel challenges Materials and Manufacturing Process June 2009 Allcock provides a clear path by first describing basic chemical principles then distinguishing between the various major materials groups and finally enriching the student by offering a variety of special examples CHOICE April 2009 Proceeding logically from the basics to materials in advanced technology it covers the fundamentals of materials chemistry including principles of materials synthesis and materials characterization methods Internationale Fachzeitschrift Metall January 2009

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complete picture of the exam pattern and level both This book is highly useful and a complete guide for the aspirants those who are willing to make future in Hotel Management

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