

Experiment: Solubility Rules and Net Ionic Equations

This is to be written up in its entirety in your laboratory notebook

OBJECTIVES

In this experiment you will mix small quantities of aqueous ionic solutions together and in some cases, observe the formation of precipitates.

OVERVIEW

1. Mix small quantities of different ionic solutions together.
2. Observe the absence or presence of a precipitate and its physical characteristics such as color, amount, etc.
3. Formulate general solubility rules based on experimental observations.
4. Write net ionic equations to describe the precipitation reactions.

SAFETY

- Wear safety goggles at all times.
- Wear shoes that cover the entire foot.
- Never touch, taste or smell chemicals.
- Keep your hands away from your face, eyes, nose and mouth.
- Always wash your hands before leaving the laboratory.

PRELAB QUESTIONS

References: Ebbing, "General Chemistry", 5th ed. pp. 91-94 (Molecular-Ionic Equations), pp. 95-98 (Precip. Reactions)

1. Write the balanced equation for dissolving the following solid ionic salts in water:



2. Write the complete ionic and the balanced net ionic equations for the following reactions:



3. Use the general solubility rules to predict which of the following compounds are soluble in water and which are not.

a. NiCl_2

b. NaOH

c. Ni(OH)_2

d. NaCl

e. $(\text{NH}_4)_2\text{S}$

f. $\text{Pb(NO}_3)_2$

g. PbS

h. NH_4NO_3

i. CaSO_4

j. K_2CO_3

k. K_2SO_4

l. CaCO_3

m. KBr

n. FeCl_3

o. FeBr_3

p. KCl

Pogil Answer Key Solubility Rules And Net Ionic Equations

Joacim Rocklöv



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Table of Contents Pogil Answer Key Solubility Rules And Net Ionic Equations

1. Understanding the eBook Pogil Answer Key Solubility Rules And Net Ionic Equations
 - The Rise of Digital Reading Pogil Answer Key Solubility Rules And Net Ionic Equations
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Answer Key Solubility Rules And Net Ionic Equations
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Answer Key Solubility Rules And Net Ionic Equations
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Answer Key Solubility Rules And Net Ionic Equations
 - Personalized Recommendations

- Pogil Answer Key Solubility Rules And Net Ionic Equations User Reviews and Ratings
- Pogil Answer Key Solubility Rules And Net Ionic Equations and Bestseller Lists
- 5. Accessing Pogil Answer Key Solubility Rules And Net Ionic Equations Free and Paid eBooks
 - Pogil Answer Key Solubility Rules And Net Ionic Equations Public Domain eBooks
 - Pogil Answer Key Solubility Rules And Net Ionic Equations eBook Subscription Services
 - Pogil Answer Key Solubility Rules And Net Ionic Equations Budget-Friendly Options
- 6. Navigating Pogil Answer Key Solubility Rules And Net Ionic Equations eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Answer Key Solubility Rules And Net Ionic Equations Compatibility with Devices
 - Pogil Answer Key Solubility Rules And Net Ionic Equations Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Answer Key Solubility Rules And Net Ionic Equations
 - Highlighting and Note-Taking Pogil Answer Key Solubility Rules And Net Ionic Equations
 - Interactive Elements Pogil Answer Key Solubility Rules And Net Ionic Equations
- 8. Staying Engaged with Pogil Answer Key Solubility Rules And Net Ionic Equations
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Answer Key Solubility Rules And Net Ionic Equations
- 9. Balancing eBooks and Physical Books Pogil Answer Key Solubility Rules And Net Ionic Equations
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Answer Key Solubility Rules And Net Ionic Equations
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Answer Key Solubility Rules And Net Ionic Equations
 - Setting Reading Goals Pogil Answer Key Solubility Rules And Net Ionic Equations
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Answer Key Solubility Rules And Net Ionic Equations
 - Fact-Checking eBook Content of Pogil Answer Key Solubility Rules And Net Ionic Equations

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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